Octane

Other names:	Oktan
	Oktanen
	Ottani
	UN 1262
	n-C8H18
	n-Octane
Inchi:	InChI=1S/C8H18/c1-3-5-7-8-6-4-2/h3-8H2,1-2H3
InchiKey:	TVMXDCGIABBOFY-UHFFFAOYSA-N
Formula:	C8H18
SMILES:	CCCCCCCC
Mol. weight [g/mol]:	114.23
CAS:	111-65-9

Physical Properties

Property code	Value	Unit	Source
af	0.3980		KDB
aigt	493.15	K	KDB
ap	343.750	K	KDB
chl	-5470.30 ± 1.60	kJ/mol	NIST Webbook
chl	-5270.40	kJ/mol	NIST Webbook
chl	-5441.30	kJ/mol	NIST Webbook
chl	-5471.80 ± 5.40	kJ/mol	NIST Webbook
chl	-5466.60	kJ/mol	NIST Webbook
chl	-5470.71 ± 0.67	kJ/mol	NIST Webbook
dm	0.00	debye	KDB
dvisc	0.0005180	Pa×s	Thermophysical Properties of Binary Mixtures of 2-Methyl-1-propanol with Hexane, Octane, and Decane at 298.15 K
dvisc	0.0005180	Pa×s	Study of densities, viscosities, and speeds of sound of binary liquid mixtures of butan-1-ol with n-alkanes (C6, C8, and C10) at T = (298.15, 303.15, and 308.15) K

dvisc	0.0005180	Pa×s	Study of molecular interactions in binary liquid mixtures of 1-octanol with n-hexane, n-octane, and n-decane using volumetric, viscometric, and acoustic properties
fll	1.00	% in Air	KDB
flu	6.50	% in Air	KDB
fpo	286.48	K	KDB
gf	16.40	kJ/mol	KDB
gyrad	4.6800		KDB
hcg	5470.71	kJ/mol	KDB
hcn	5074.565	kJ/mol	KDB
hf	-208.40 ± 0.67	kJ/mol	NIST Webbook
hf	-208.60	kJ/mol	KDB
hf	-208.70	kJ/mol	NIST Webbook
hfl	-250.30 ± 1.80	kJ/mol	NIST Webbook
hfl	-250.00 ± 0.84	kJ/mol	NIST Webbook
hfus	16.48	kJ/mol	Joback Method
hvap	41.48	kJ/mol	NIST Webbook
hvap	41.60 ± 0.08	kJ/mol	NIST Webbook
hvap	41.60 ± 0.08	kJ/mol	NIST Webbook
hvap	34.00	kJ/mol	NIST Webbook
hvap	41.50	kJ/mol	NIST Webbook
hvap	41.50 ± 0.10	kJ/mol	NIST Webbook
hvap	41.53	kJ/mol	NIST Webbook
hvap	41.50 ± 0.10	kJ/mol	NIST Webbook
hvap	41.50 ± 0.10	kJ/mol	NIST Webbook
hvap	41.60	kJ/mol	NIST Webbook
ie	10.25	eV	NIST Webbook
ie	9.71 ± 0.15	eV	NIST Webbook
ie	9.80 ± 0.15	eV	NIST Webbook
ie	9.79	eV	NIST Webbook
ie	9.80 ± 0.10	eV	NIST Webbook
ie	10.01	eV	NIST Webbook
log10ws	-5.24		Aqueous Solubility Prediction Method
log10ws	-5.24		Estimated Solubility Method
logp	3.367		Crippen Method
mcvol	123.580	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
pc	2487.00	kPa	NIST Webbook
pc	2486.93 ± 17.23	kPa	NIST Webbook
pc	2490.00	kPa	KDB
pc	2490.00 ± 30.00	kPa	NIST Webbook

pc	2495.00 ± 20.00	kPa	NIST Webbook
pc	2470.00 ± 98.07	kPa	NIST Webbook
pc	2480.00 ± 5.00	kPa	NIST Webbook
pc	2490.00 ± 30.00	kPa	NIST Webbook
pc	2497.12 ± 19.99	kPa	NIST Webbook
pc	2486.20 ± 40.53	kPa	NIST Webbook
pc	2488.00 ± 40.53	kPa	NIST Webbook
pc	2492.59 ± 50.66	kPa	NIST Webbook
pc	2497.10 ± 20.00	kPa	NIST Webbook
sg	467.06 ± 0.92	J/mol×K	NIST Webbook
sl	359.80	J/mol×K	NIST Webbook
sl	361.20	J/mol×K	NIST Webbook
sl	359.80	J/mol×K	NIST Webbook
tb	399.00 ± 0.60	K	NIST Webbook
tb	398.15 ± 1.00	K	NIST Webbook
tb	397.65 ± 0.30	K	NIST Webbook
tb	399.00 ± 0.70	K	NIST Webbook
tb	399.00 ± 1.00	K	NIST Webbook
tb	398.83	K	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
tb	398.95 ± 0.20	K	NIST Webbook
tb	397.75 ± 0.50	K	NIST Webbook
tb	398.74 ± 0.15	K	NIST Webbook
tb	398.55 ± 0.50	K	NIST Webbook
tb	398.35 ± 0.30	K	NIST Webbook
tb	398.74 ± 0.10	K	NIST Webbook
tb	398.75 ± 0.50	K	NIST Webbook
tb	398.55 ± 0.50	K	NIST Webbook
tb	396.00 ± 2.00	K	NIST Webbook
tb	398.15 ± 1.00	K	NIST Webbook
tb	398.95 ± 0.30	K	NIST Webbook
tb	398.15 ± 2.00	K	NIST Webbook
tb	398.95 ± 0.30	K	NIST Webbook
tb	398.64 ± 0.10	K	NIST Webbook
tb	398.45 ± 0.50	K	NIST Webbook
tb	398.82	K	KDB
tb	398.57	K	Isobaric vapor-liquid equilibrium for binary systemof 2-ethylthiophene + n-octane at 101.33 kPa

tb	398.83	K	Solutions of alkyl methanoates and alkanes: Simultaneous modeling of phase equilibria and mixing properties. Estimation of behavior by UNIFAC with recalculation of parameters
tb	398.79	K	Isothermal and isobaric (vapour + liquid) equilibria of (alpha-pinene + n-butanol + n-octane)
tb	398.60	K	Excess molar volumes of the ternary system {methylcyclohexane (1) + cyclohexane (2) + n-alkanes (3)} at T = 298.15 K
tb	399.15	K	Excess volumes, densities, speeds of sound and viscosities for the binary systems of diisopropyl ether with hydrocarbons at 303.15K
tb	398.65	K	Densities and Excess Molar Properties of Dimethyl Carbonate with Alkanes (C6 to C10) and VLE of Dimethyl Carbonate with Alkanes (C9 to C10) at 101.3 kPa
tb	396.70 ± 1.50	K	NIST Webbook
tb	398.49	K	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
tb	398.74	K	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
tb	398.80 ± 0.20	K	NIST Webbook
tb	398.80 ± 0.20	K	NIST Webbook
tb	398.90	K	NIST Webbook
tb	398.73 ± 0.10	K	NIST Webbook
tb	398.80	K	NIST Webbook
tb	398.76 ± 0.20	K	NIST Webbook
tb	398.96 ± 0.20	K	NIST Webbook
tb	398.96 ± 0.50	K	NIST Webbook
tb	398.93 ± 0.08	K	NIST Webbook
tb	398.70 ± 0.50	K	NIST Webbook
tb	398.75 ± 0.15	K	NIST Webbook
tb	398.55 ± 0.20	K	NIST Webbook
tb	398.12 ± 0.20	K	NIST Webbook
tb	398.77 ± 0.10	K	NIST Webbook

tb	398.95 ± 0.40	K	NIST Webbook
tb	398.76 ± 0.20	K	NIST Webbook
tb	398.75 ± 0.50	K	NIST Webbook
tb	398.55 ± 0.30	K	NIST Webbook
tb	398.77 ± 0.15	K	NIST Webbook
tb	398.27 ± 0.30	K	NIST Webbook
tb	398.15 ± 0.60	K	NIST Webbook
tb	398.40 ± 1.00	K	NIST Webbook
tb	398.85 ± 0.25	K	NIST Webbook
tb	398.82 ± 0.30	K	NIST Webbook
tb	399.15 ± 1.50	K	NIST Webbook
tb	397.80 ± 0.30	K	NIST Webbook
tb	396.40 ± 2.00	K	NIST Webbook
tb	398.85 ± 0.40	K	NIST Webbook
tb	399.09 ± 0.30	K	NIST Webbook
tb	398.75 ± 0.10	K	NIST Webbook
tb	398.81 ± 0.44	K	NIST Webbook
tb	398.55 ± 0.30	K	NIST Webbook
tb	398.71 ± 0.20	K	NIST Webbook
tb	398.80 ± 0.40	K	NIST Webbook
tb	398.95 ± 0.50	K	NIST Webbook
tb	393.15 ± 3.00	K	NIST Webbook
tb	398.81 ± 0.10	K	NIST Webbook
tb	398.90 ± 0.20	K	NIST Webbook
tb	398.25 ± 0.30	K	NIST Webbook
tb	398.62 ± 0.30	K	NIST Webbook
tb	397.85 ± 0.30	K	NIST Webbook
tb	397.20 ± 2.00	K	NIST Webbook
tb	398.75 ± 0.30	K	NIST Webbook
tb	398.55 ± 0.50	K	NIST Webbook
tb	398.85 ± 0.50	K	NIST Webbook
tb	397.35 ± 0.50	K	NIST Webbook
tb	398.75 ± 0.30	K	NIST Webbook
tb	398.90 ± 0.30	K	NIST Webbook
tb	398.75 ± 0.40	K	NIST Webbook
tb	398.81 ± 0.01	K	NIST Webbook
tb	397.65 ± 0.50	K	NIST Webbook
tb	397.40 ± 1.00	K	NIST Webbook
tb	398.15 ± 2.00	K	NIST Webbook
tb	398.70 ± 0.20	K	NIST Webbook
tb	398.95 ± 0.50	K	NIST Webbook
tb	398.90 ± 0.20	K	NIST Webbook
tb	398.75 ± 0.30	K	NIST Webbook
tb	398.81 ± 0.10	K	NIST Webbook

tb	398.80 ± 0.30	K	NIST Webbook
tb	397.15 ± 1.00	K	NIST Webbook
tb	398.75 ± 0.20	K	NIST Webbook
tb	401.65 ± 0.50	K	NIST Webbook
tb	398.75 ± 0.30	K	NIST Webbook
tb	398.90 ± 0.20	K	NIST Webbook
tb	398.81 ± 0.20	K	NIST Webbook
tb	398.95 ± 0.20	K	NIST Webbook
tb	398.80 ± 0.05	K	NIST Webbook
tb	398.80 ± 0.40	K	NIST Webbook
tc	568.92	K	Determination of the Critical Properties of C6 C10 n-Alkanes and Their Binary Systems Using a Flow Apparatus
tc	568.80	K	The Critical Temperatures of a Number of (i) (Chloroalkane (C3 C4) + Hydrocarbon (C6 C7)) Binary Mixtures and (ii) (Aromatic Halocarbon (Chlorobenzene, Fluorobenzene, 1,2-Dichlorobenzene, or 1,3-Dichlorobenzene) + Alkane (C8)) Binary Mixtures
tc	568.90	K	Measurement of Critical Properties for Binary and Ternary Mixtures Containing n-Butanol and n-Alkane
tc	568.92	K	Experimental determination of critical data of multi-component mixtures containing potential gasoline additives 2-butanol by a flow-type apparatus
tc	568.90	K	Measurement of critical properties for binary and ternary mixtures containing potential gasoline additive diethyl carbonate (DEC)
tc	568.70	K	KDB
tf	216.26	K	Aqueous Solubility Prediction Method
tf	216.30	K	KDB
tt	216.37	K	KDB
tt	216.37 ± 0.40	K	NIST Webbook
tt	216.37 ± 0.03	K	NIST Webbook
tt	216.36 ± 0.20	K	NIST Webbook
tt	216.37 ± 0.20	K	NIST Webbook
tt	216.36 ± 0.05	K	NIST Webbook

tt	216.37 ± 0.02	K	NIST Webbook
tt	215.80 ± 0.20	K	NIST Webbook
tt	215.85 ± 0.20	K	NIST Webbook
tt	215.60 ± 0.25	K	NIST Webbook
vc	0.492	m3/kmol	KDB
vc	0.492	m3/kmol	NIST Webbook
zc	0.2590860		KDB
zra	0.26		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	238.95 ± 0.48	J/mol×K	398.15	NIST Webbook
cpg	242.67	J/mol×K	405.70	NIST Webbook
cpg	250.59 ± 0.50	J/mol×K	423.15	NIST Webbook
cpg	263.02 ± 0.53	J/mol×K	448.15	NIST Webbook
cpg	297.00 ± 0.59	J/mol×K	523.15	NIST Webbook
cpg	274.84 ± 0.55	J/mol×K	473.15	NIST Webbook
cpg	285.98 ± 0.57	J/mol×K	498.15	NIST Webbook
cpg	295.39	J/mol×K	522.70	NIST Webbook
cpg	270.70	J/mol×K	462.50	NIST Webbook
cpg	232.74 ± 0.47	J/mol×K	385.65	NIST Webbook
cpl	253.94	J/mol×K	298.15	Thermodynamic behaviour of the binary systems dimethyl carbonate + n-octane or n-nonane
cpl	254.14	J/mol×K	298.15	NIST Webbook
cpl	253.20	J/mol×K	298.00	NIST Webbook
cpl	252.92	J/mol×K	298.15	NIST Webbook
cpl	254.07	J/mol×K	298.15	NIST Webbook
cpl	252.40	J/mol×K	298.00	NIST Webbook
cpl	254.02	J/mol×K	298.15	NIST Webbook
cpl	252.40	J/mol×K	297.54	NIST Webbook
cpl	253.72	J/mol×K	298.15	NIST Webbook
cpl	254.18	J/mol×K	298.15	NIST Webbook
cpl	255.68	J/mol×K	298.15	NIST Webbook
cpl	254.11	J/mol×K	298.15	NIST Webbook
cpl	255.68	J/mol×K	298.15	NIST Webbook
cpl	255.68	J/mol×K	298.15	NIST Webbook
cpl	255.68	J/mol×K	298.15	NIST Webbook

cpl	262.20	J/mol×K	318.15	NIST Webbook
cpl	254.70	J/mol×K	299.00	NIST Webbook
cpl	269.31	J/mol×K	333.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions
cpl	267.24	J/mol×K	328.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions
cpl	264.78	J/mol×K	323.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions
cpl	262.70	J/mol×K	318.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions
cpl	260.87	J/mol×K	313.16	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions

cpl	258.05	J/mol×K	308.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions	
cpl	253.89	J/mol×K	298.15	NIST Webbook	
cpl	256.50	J/mol×K	303.16	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions	
cpl	253.98	J/mol×K	298.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions	
cpl	252.24	J/mol×K	293.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions	
cpl	250.10	J/mol×K	288.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions	
cpl	253.93	J/mol×K	299.80	NIST Webbook	

cpl	258.03	J/mol×K	308.15	Thermodynamic behaviour of the binary systems dimethyl carbonate + n-octane or n-nonane
cpl	251.50	J/mol×K	298.30	NIST Webbook
cpl	250.12	J/mol×K	288.15	Thermodynamic behaviour of the binary systems dimethyl carbonate + n-octane or n-nonane
cpl	247.70	J/mol×K	293.70	NIST Webbook
cpl	248.16	J/mol×K	283.15	Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine Water} and {N-Methylpiperidine Water} Systems of Low to Moderate Amine Compositions
dvisc	0.0004720	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0004500	Pa×s	308.15	Density, Viscosity, and Speed of Sound of (Methyl Benzoate + Cyclohexane), (Methyl Benzoate + n-Hexane), (Methyl Benzoate + Heptane), and (Methyl Benzoate + Octane) at Temperatures of (303.15, 308.15, and 313.15) K

dvisc	0.0004910	Pa×s	303.15	Density, Viscosity, and Speed of Sound of (Methyl Benzoate + Cyclohexane), (Methyl Benzoate + n-Hexane), (Methyl Benzoate + Heptane), and (Methyl Benzoate + Octane) at Temperatures of (303.15, 308.15, and 313.15) K
dvisc	0.0005087	Pa×s	298.15	Viscosity and Density of Binary Mixtures of Ethyl Alcohol with n-Alkanes (C6, C8, and C10)
dvisc	0.0005374	Pa×s	293.15	Viscosity and Density of Binary Mixtures of Ethyl Alcohol with n-Alkanes (C6, C8, and C10)
dvisc	0.0004358	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of JP-10 with n-Octane or n-Decane at Several Temperatures
dvisc	0.0004795	Pa×s	303.15	Densities and Viscosities of Binary Mixtures of JP-10 with n-Octane or n-Decane at Several Temperatures
dvisc	0.0005177	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of JP-10 with n-Octane or n-Decane at Several Temperatures
dvisc	0.0005439	Pa×s	293.15	Densities and Viscosities of Binary Mixtures of JP-10 with n-Octane or n-Decane at Several Temperatures

dvisc	0.0003570	Pa×s	333.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	
dvisc	0.0003750	Pa×s	328.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	
dvisc	0.0004426	Pa×s	313.30	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa	
dvisc	0.0003940	Pa×s	323.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	
dvisc	0.0004150	Pa×s	318.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	
dvisc	0.0004340	Pa×s	313.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	
dvisc	0.0004590	Pa×s	308.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	
dvisc	0.0004860	Pa×s	303.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K	

dvisc	0.0005140	Pa×s	298.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K
dvisc	0.0005450	Pa×s	293.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K
dvisc	0.0005840	Pa×s	288.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K
dvisc	0.0006230	Pa×s	283.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K
dvisc	0.0006710	Pa×s	278.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K
dvisc	0.0007180	Pa×s	273.15	Densities and Viscosities of MTBE + Heptane or Octane at p) 0.1 MPa from (273.15 to 363.15) K
dvisc	0.0002840	Pa×s	353.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures
dvisc	0.0003139	Pa×s	343.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures

dvisc	0.0003495	Pa×s	333.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures
dvisc	0.0003864	Pa×s	323.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures
dvisc	0.0004306	Pa×s	313.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures
dvisc	0.0004837	Pa×s	303.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures
dvisc	0.0005206	Pa×s	298.15	Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures
dvisc	0.0002610	Pa×s	363.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure

dvisc	0.0002870	Pa×s	353.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0003140	Pa×s	343.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0003430	Pa×s	333.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0003800	Pa×s	323.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0004220	Pa×s	313.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure

dvisc	0.0004150	Pa×s	313.15	Density, Viscosity, and Speed of Sound of (Methyl Benzoate + Cyclohexane), (Methyl Benzoate + n-Hexane), (Methyl Benzoate + Heptane), and (Methyl Benzoate + Octane) at Temperatures of (303.15, 308.15, and 313.15) K
dvisc	0.0005010	Pa×s	298.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0004300	Pa×s	313.15	A Study on Properties Derived from Densities and Viscosities for the Ternary Systems (Methyl Pentanoate or Methyl Heptanoate) + n-Octane + 1-Hexanol and their Binary Subsystems at Various Temperatures.
dvisc	0.0004780	Pa×s	303.15	Dynamic Viscosities of 2-Pentanol with Alkanes (Octane, Decane, and Dodecane) at Three Temperatures T) (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO Interaction Parameters

dvisc	0.0005060	Pa×s	298.15	Dynamic Viscosities of 2-Pentanol with Alkanes (Octane, Decane, and Dodecane) at Three Temperatures T) (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO Interaction Parameters
dvisc	0.0005380	Pa×s	293.15	Dynamic viscosities of 2-butanol with alkanes (C8, C10, and C12) at several temperatures
dvisc	0.0005060	Pa×s	298.15	Dynamic viscosities of 2-butanol with alkanes (C8, C10, and C12) at several temperatures
dvisc	0.0004780	Pa×s	303.15	Dynamic viscosities of 2-butanol with alkanes (C8, C10, and C12) at several temperatures
dvisc	0.0004911	Pa×s	303.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0004414	Pa×s	313.30	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer

dvisc	0.0003967	Pa×s	323.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0003632	Pa×s	333.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0003352	Pa×s	343.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0003012	Pa×s	353.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0002766	Pa×s	363.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0002545	Pa×s	373.30	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer

dvisc	0.0002341	Pa×s	383.00	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0002208	Pa×s	393.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0002021	Pa×s	403.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001924	Pa×s	413.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001762	Pa×s	423.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001642	Pa×s	433.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer

dvisc	0.0001527	Pa×s	443.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001416	Pa×s	453.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001322	Pa×s	463.20	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001228	Pa×s	473.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001139	Pa×s	483.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0001069	Pa×s	493.10	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer

dvisc	0.0000995	Pa×s	503.00	Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer
dvisc	0.0004913	Pa×s	303.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0005380	Pa×s	293.15	Dynamic Viscosities of 2-Pentanol with Alkanes (Octane, Decane, and Dodecane) at Three Temperatures T) (293.15, 298.15, and 303.15) K. New UNIFAC-VISCO Interaction Parameters
dvisc	0.0003961	Pa×s	323.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0003634	Pa×s	333.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0003311	Pa×s	343.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa

dvisc	0.0003013	Pa×s	353.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0002765	Pa×s	363.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0002540	Pa×s	373.30	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0002340	Pa×s	383.00	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0002205	Pa×s	393.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0002019	Pa×s	403.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001905	Pa×s	413.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa

dvisc	0.0001761	Pa×s	423.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001643	Pa×s	433.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001526	Pa×s	443.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001417	Pa×s	453.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001321	Pa×s	463.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001228	Pa×s	473.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0001138	Pa×s	483.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa

dvisc	0.0001068	Pa×s	493.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000993	Pa×s	503.00	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000901	Pa×s	513.20	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000825	Pa×s	523.30	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000753	Pa×s	533.40	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000678	Pa×s	543.40	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000607	Pa×s	553.40	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa

dvisc	0.0005340	Pa×s	293.15	Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure
dvisc	0.0000444	Pa×s	573.30	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0000325	Pa×s	583.10	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
dvisc	0.0005014	Pa×s	303.20	Experimental Measurement of JP-10 Viscosity at 242.7-753.3 K under Pressures up to 6.00 MPa
dvisc	0.0003078	Pa×s	353.10	Experimental Measurement of JP-10 Viscosity at 242.7-753.3 K under Pressures up to 6.00 MPa
dvisc	0.0002070	Pa×s	403.20	Experimental Measurement of JP-10 Viscosity at 242.7-753.3 K under Pressures up to 6.00 MPa
dvisc	0.0001465	Pa×s	453.10	Experimental Measurement of JP-10 Viscosity at 242.7-753.3 K under Pressures up to 6.00 MPa
dvisc	0.0001049	Pa×s	503.00	Experimental Measurement of JP-10 Viscosity at 242.7-753.3 K under Pressures up to 6.00 MPa

dvisc	0.0000692	Pa×s	553.40	Experimental Measurement of JP-10 Viscosity at 242.7-753.3 K under Pressures up to 6.00 MPa
dvisc	0.0006247	Pa×s	283.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering
dvisc	0.0005523	Pa×s	293.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering
dvisc	0.0004828	Pa×s	303.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering
dvisc	0.0004320	Pa×s	313.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering
dvisc	0.0003870	Pa×s	323.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering

dvisc	0.0003079	Pa×s	348.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0002487	Pa×s	373.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0002028	Pa×s	398.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0001665	Pa×s	423.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0001393	Pa×s	448.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0001168	Pa×s	473.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	

dvisc	0.0000947	Pa×s	498.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0000755	Pa×s	523.15	Liquid Viscosity and Surface Tension of n-Hexane, n-Octane, n-Decane, and n-Hexadecane up to 573 K by Surface Light Scattering	
dvisc	0.0006140	Pa×s	283.15	A Study on Properties Derived from Densities and Viscosities for the Ternary Systems (Methyl Pentanoate or Methyl Heptanoate) + n-Octane + 1-Hexanol and their Binary Subsystems at Various Temperatures.	
dvisc	0.0005420	Pa×s	293.15	A Study on Properties Derived from Densities and Viscosities for the Ternary Systems (Methyl Pentanoate or Methyl Heptanoate) + n-Octane + 1-Hexanol and their Binary Subsystems at Various Temperatures.	

dvisc	0.0005080	Pa×s	298.15	A Study on Properties Derived from Densities and Viscosities for the Ternary Systems (Methyl Pentanoate or Methyl Heptanoate) + n-Octane + 1-Hexanol and their Binary Subsystems at Various Temperatures.
dvisc	0.0004830	Pa×s	303.15	A Study on Properties Derived from Densities and Viscosities for the Ternary Systems (Methyl Pentanoate or Methyl Heptanoate) + n-Octane + 1-Hexanol and their Binary Subsystems at Various Temperatures.
dvisc	0.0000532	Pa×s	563.30	Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa
hfust	20.09	kJ/mol	215.60	NIST Webbook
hfust	20.65	kJ/mol	215.80	NIST Webbook
hfust	20.74	kJ/mol	216.40	NIST Webbook
hfust	21.80	kJ/mol	216.60	NIST Webbook
hfust	20.74	kJ/mol	216.38	NIST Webbook
hsubt	68.10	kJ/mol	216.00	NIST Webbook
hvapt	39.10 ± 0.10	kJ/mol	333.00	NIST Webbook
hvapt	35.40 ± 0.10	kJ/mol	344.00	NIST Webbook
hvapt	36.70 ± 0.10	kJ/mol	328.00	NIST Webbook
hvapt	38.00 ± 0.10	kJ/mol	311.00	NIST Webbook
hvapt	43.00	kJ/mol	257.00	NIST Webbook
hvapt	37.80 ± 0.10	kJ/mol	353.00	NIST Webbook
hvapt	39.20	kJ/mol	363.00	NIST Webbook
hvapt	40.50 ± 0.10	kJ/mol	313.00	NIST Webbook
hvapt	41.90	kJ/mol	315.50	NIST Webbook
hvapt	41.20	kJ/mol	348.50	NIST Webbook

hvapt	34.90	kJ/mol	537.50	NIST Webbook
hvapt	35.50	kJ/mol	469.00	NIST Webbook
hvapt	36.30	kJ/mol	414.00	NIST Webbook
hvapt	44.40	kJ/mol	247.00	NIST Webbook
hvapt	39.40	kJ/mol	443.00	NIST Webbook
hvapt	34.41	kJ/mol	398.80	NIST Webbook
hvapt	41.75	kJ/mol	298.00	Enthalpies of Vaporization and Vapor Pressures of Some Deuterated Hydrocarbons. Liquid-Vapor Pressure Isotope Effects
hvapt	34.41	kJ/mol	398.80	KDB
hvapt	41.00	kJ/mol	348.50	NIST Webbook
pvap	60.63	kPa	381.34	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	121.60	kPa	405.51	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	119.25	kPa	404.79	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	117.10	kPa	404.10	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	116.17	kPa	403.81	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	113.09	kPa	402.81	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	110.71	kPa	402.03	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	108.30	kPa	401.23	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	106.23	kPa	400.53	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	103.61	kPa	399.61	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	101.32	kPa	398.83	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	100.68	kPa	398.59	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	99.59	kPa	398.19	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	98.73	kPa	397.90	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	96.14	kPa	396.95	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	93.64	kPa	396.04	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	124.10	kPa	406.27	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	91.38	kPa	395.14	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	89.39	kPa	394.37	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	87.09	kPa	393.47	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	84.94	kPa	392.60	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	82.51	kPa	391.60	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	80.05	kPa	390.56	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	77.94	kPa	389.65	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	75.43	kPa	388.55	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	73.02	kPa	387.45	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	70.52	kPa	386.29	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	65.07	kPa	383.64	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	62.68	kPa	382.42	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	126.61	kPa	407.03	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	58.20	kPa	380.03	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	55.02	kPa	378.25	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	52.90	kPa	377.00	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	50.59	kPa	375.61	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	48.04	kPa	374.01	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	45.57	kPa	372.39	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	43.19	kPa	370.77	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	40.82	kPa	369.09	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	38.88	kPa	367.65	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	36.53	kPa	365.82	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	34.16	kPa	363.86	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	31.58	kPa	361.59	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	28.55	kPa	358.82	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	26.70	kPa	356.97	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	25.56	kPa	355.76	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	23.14	kPa	353.05	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	21.09	kPa	350.64	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	60.00	kPa	380.84	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	60.00	kPa	380.78	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	60.00	kPa	380.87	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	60.00	kPa	380.76	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	60.00	kPa	380.85	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	60.00	kPa	380.94	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	129.50	kPa	407.89	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	132.46	kPa	408.75	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	135.53	kPa	409.62	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	137.42	kPa	410.15	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	67.34	kPa	384.76	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	139.87	kPa	410.83	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	185.56	kPa	422.10	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	188.02	kPa	422.64	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	190.85	kPa	423.26	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	193.15	kPa	423.76	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	195.02	kPa	424.15	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	197.87	kPa	424.76	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	199.97	kPa	425.20	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	30.00	kPa	359.76	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	60.00	kPa	380.43	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	90.00	kPa	394.09	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	101.30	kPa	398.49	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	30.00	kPa	360.06	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	60.00	kPa	380.94	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	90.00	kPa	394.60	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System
pvap	101.30	kPa	398.74	Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Sulfuric Acid + Water + Ethanol System

pvap	40.00	kPa	368.32	Vapor Liquid Equilibria Measurements for the Nine n-Alkane/Ketone Pairs Comprising 2-, 3-, and 4-Heptanone with n-Octane, n-Nonane, and n-Decane
pvap	40.00	kPa	367.97	Vapor Liquid Equilibria Measurements for the Nine n-Alkane/Ketone Pairs Comprising 2-, 3-, and 4-Heptanone with n-Octane, n-Nonane, and n-Decane
pvap	40.00	kPa	368.30	Vapor Liquid Equilibria Measurements for the Nine n-Alkane/Ketone Pairs Comprising 2-, 3-, and 4-Heptanone with n-Octane, n-Nonane, and n-Decane
pvap	60.00	kPa	380.82	Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane
pvap	101.33	kPa	398.57	Isobaric vapor-liquid equilibrium for binary systemof 2-ethylthiophene + n-octane at 101.33 kPa
pvap	23.16	kPa	353.15	Measurement of vapor.liquid equilibria (VLE) and excess enthalpies (HE) of binary systems with 1-alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and prediction of these properties and A using modified UNIFAC (Dortmund)

pvap	33.91	kPa	363.15	An equipment for dynamic measurements of vapour liquid equilibria and results in binary systems containing cyclohexylamine	
pvap	23.61	kPa	353.15	An equipment for dynamic measurements of vapour liquid equilibria and results in binary systems containing cyclohexylamine	
pvap	16.04	kPa	343.15	An equipment for dynamic measurements of vapour liquid equilibria and results in binary systems containing cyclohexylamine	
pvap	10.59	kPa	333.15	An equipment for dynamic measurements of vapour liquid equilibria and results in binary systems containing cyclohexylamine	
pvap	114.65	kPa	403.32	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
rfi	1.39540		298.15	Solubilities of Bis (2,2,6,6-Tetramethyl-4-Piperidinyl) Maleate in Hexane, Heptane, Octane, m-Xylene and Tetrahydrofuran from (253.15 to 310.15) K	

rfi	1.39780	298.10	Excess Enthalpies and Thermal Conductivity Coefficients for Binary Mixtures of Carbon Tetrachloride and Four Alkanes (C5 to C8) at a Temperature of 298.15 K
rfi	1.39650	293.15	Isobaric Vapor-Liquid Equilibria of Hexane + 1-Decene and Octane + 1-Decene Mixtures
rfi	1.38550	318.15	Multiproperty Correlation of Experimental Data of the Binaries Propyl Ethanoate + Alkanes (Pentane to Decane). New Experimental Information for Vapor Liquid Equilibrium and Mixing Properties
rfi	1.39520	298.15	Multiproperty Correlation of Experimental Data of the Binaries Propyl Ethanoate + Alkanes (Pentane to Decane). New Experimental Information for Vapor Liquid Equilibrium and Mixing Properties
rfi	1.39230	303.15	Densities, Viscosities, and Refractive Indices of Binary Mixtures of 1,2,3,4-Tetrahydronaphthalene with Some n-Alkanes at T = (293.15 to 313.15) K

rfi	1.39700	293.15	Densities, Viscosities, and Refractive Indices of Binary Mixtures of 1,2,3,4-Tetrahydronaphthalene with Some n-Alkanes at T = (293.15 to 313.15) K
rfi	1.39519	298.15	Extraction of Benzene from Aliphatic Compounds Using Commercial Ionic Liquids as Solvents: Study of the Liquid Liquid Equilibrium at T = 298.15 K
rfi	1.39519	298.15	Liquid-Liquid Equilibria of the Ternary Systems of Alkane + Aromatic + 1-Ethylpyridinium Ethylsulfate Ionic Liquid at T = (283.15 and 298.15) K
rfi	1.39519	298.15	Separation of Benzene from Linear Alkanes (C6-C9) Using 1-Ethyl-3-Methylimidazolium Ethylsulfate at T = 298.15 K
rfi	1.39460	298.15	Ternary Liquid-Liquid(-Liquid) Equilibria of Aniline + Cyclohexylamine + Water, Aniline + Cyclohexylamine + Octane, Aniline + Water + Toluene, and Aniline + Water + Octane

rfi	1.39460	298.15	Liquid-Liquid(-Liquid) Equilibria in Ternary Systems of Water + Cyclohexylamine + Aromatic Hydrocarbon (Toluene or Propylbenzene) or Aliphatic Hydrocarbon (Heptane or Octane)
rfi	1.39730	293.15	Vapor-Liquid Equilibria for the Binary Mixture alpha-Pinene + Octane
rfi	1.39750	293.15	Infinite Dilution Activity Coefficients of Hydrocarbons in Triethylene Glycol and Tetraethylene Glycol
rfi	1.39550	298.15	Measurement and Prediction of Excess Properties of Binary Mixtures Methyl Decanoate + an Even-Numbered n-Alkane (C6-C16) at 298.15 K
rfi	1.39530	298.15	Isothermal Bubble Pressure Data for the Binary System of C2F6 and n-Octane
rfi	1.39780	293.15	Isothermal Bubble Pressure Data for the Binary System of C2F6 and n-Octane
rfi	1.39800	298.15	P-rho-T Data and Modeling for Propan-1-ol + n-Octane or n-Nonane or n-Decane from 313.15 K to 363.15 K and 1 MPa to 20 MPa

rfi	1.39510	298.15	Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K.
rfi	1.39520	298.15	A Study on Alkane + Ester + Ester Systems. Physicochemical Behavior of Binaries and Ternaries of Octane or Iso-octane with Methyl Esters (Ethanoate, Butanoate, Pentanoate)
rfi	1.39780	293.15	Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K
rfi	1.39519	298.15	Extraction of toluene from aliphatic compounds using an ionic liquid as solvent: Influence of the alkane on the (liquid + liquid) equilibrium
rfi	1.40070	284.73	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.40120	283.88	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region

rfi	1.40170	282.65	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.40210	281.66	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.40270	280.51	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.40280	280.36	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.40320	279.68	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.40400	277.80	The measurements of coexistence curves and light scattering for {xC6H5CN + (1-x)CH3(CH2)6CH3} in the critical region
rfi	1.39500	298.15	p, Vm, T) measurements of (octane + benzene) at temperatures from (298.15 to 328.15) K and at pressures up to 40 MPa

rfi	1.39500	298.15	Excess molar volumes of (octane + benzene, or toluene, or 1,3-xylene, or 1,3,5-trimethylbenzene) at temperatures between (298.15 and 328.15) K
rfi	1.39500	298.15	(P,Vm,T) Measurements of (octane + 1-chlorohexane) at temperatures from 298.15 K to 328.15 K and at pressures up to 40 MPa
rfi	1.38880	308.15	Excess properties of the binary mixtures of methylcyclohexane + alkanes (C6 to C12) at T = 298.15 K to T = 308.15 K
rfi	1.39160	303.15	Excess properties of the binary mixtures of methylcyclohexane + alkanes (C6 to C12) at T = 298.15 K to T = 308.15 K
rfi	1.39260	303.15	Physical properties of {anisole + n-alkanes} at temperatures between (293.15 and 303.15) K
rfi	1.39470	298.15	Physical properties of {anisole + n-alkanes} at temperatures between (293.15 and 303.15) K
rfi	1.39650	293.15	Physical properties of {anisole + n-alkanes} at temperatures between (293.15 and 303.15) K
rfi	1.39538	298.15	Liquid liquid equilibria of lactam containing binary systems

rfi	1.39460	298.15	Vapor liquid equilibria in ternary systems of associating components (water, aniline, cyclohexylamine) and hydrocarbons (octane or toluene)
rfi	1.39504	298.15	Separation of aromatic hydrocarbons from alkanes using ammonium ionic liquid C2NTf2 at T = 298.15K
rfi	1.53910	298.15	(Vapor + liquid) equilibrium of binary mixtures formed by N,N-dimethyl formamide with some compounds at 95.1 kPa
rfi	1.39460	298.15	Activity Coefficients at Infinite Dilution of Cylcohexylamine + Octane, Toluene, Ethylbenzene, or Aniline and Excess Molar Volumes in Binary Mixtures of Cyclohexylamine + Heptane, Octane, Nonane, Decane, Undecane, Aniline, or Water
rfi	1.39505	298.15	KDB
rfi	1.39470	298.15	Excess properties of the binary mixtures of methylcyclohexane + alkanes (C6 to C12) at T = 298.15 K to T = 308.15 K
rfi	1.39730	293.15	Limiting Activity Coefficients by Comparative Tensimetry: 1-Propanol and 1-Butanol in Heptane and in Octane

rhol	698.71	kg/m3	298.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	686.36	kg/m3	313.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	682.26	kg/m3	318.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	678.13	kg/m3	323.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	706.54	kg/m3	288.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	702.53	kg/m3	293.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa

rhol	698.51	kg/m3	298.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	694.48	kg/m3	303.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	690.42	kg/m3	308.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	686.34	kg/m3	313.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	682.24	kg/m3	318.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	690.44	kg/m3	308.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	673.95	kg/m3	328.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	669.77	kg/m3	333.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa

rhol	665.55	kg/m3	338.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	661.28	kg/m3	343.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	703.00	kg/m3	293.15	Experimental Solubility Data for Binary Mixtures of Ethane and 2,2,4-Trimethylpentane at Pressures up to 6 MPa Using a New Variable-Volume Sapphire Cell
rhol	698.71	kg/m3	298.15	Liquid-Liquid Equilibria for 2-Phenylethan-1-ol + Alkane Systems
rhol	698.70	kg/m3	298.15	Liquid-Liquid Equilibrium Data for Ternary Systems Containing Alkanes (n-Pentane, n-Hexane, n-Heptane, and n-Octane) + Alcohol (Methanol and Ethanol) + Protic Ionic Liquid (2-HEAF)
rhol	698.50	kg/m3	298.15	Viscosities of Dimethyl Carbonate or Diethyl Carbonate with Alkanes at Four Temperatures. New UNIFAC-VISCO Parameters

rhol	702.87	kg/m3	293.15	Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol + Acetone, + Octane, + Heptane at 323.15
rhol	698.72	kg/m3	298.15	Excess Enthalpies of the Ternary Mixtures: Diisopropyl Ether + 3-Methylpentane + (Octane or Decane) at 298.15 K
rhol	698.72	kg/m3	298.15	Excess Enthalpies of Binary Mixtures of 1-Hexene with Some n-Alkanes at 298.15 K
rhol	698.55	kg/m3	298.15	Thermodynamics of Mixtures Containing Ethers. Part III. Liquid-Liquid Equilibria for 2,5,8,11-Tetraoxadodecane or 2,5,8,11,14-Pentaoxapentadecane + Selected N-Alkanes
rhol	698.64	kg/m3	298.15	Thermodynamics of Mixtures Containing a Strongly Polar Compound. 8. Liquid-Liquid Equilibria for N,N-Dialkylamide + Selected N-Alkanes
rhol	710.76	kg/m3	283.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	706.76	kg/m3	288.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}

rhol	702.74	kg/m3	293.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	694.49	kg/m3	303.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	694.66	kg/m3	303.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	690.58	kg/m3	308.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	686.49	kg/m3	313.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	682.37	kg/m3	318.15	Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 -x) (Octane or Nonane)}
rhol	698.49	kg/m3	298.15	Liquid-Liquid(-Liquid) Equilibria in Ternary Systems of Aliphatic Hydrocarbons (Heptane or Octane) + Phenols + Water

rhol	694.52	kg/m3	303.15	Vapor-Liquid Equilibria in Ternary Systems of Toluene or Octane + Phenols + Water
rhol	698.57	kg/m3	298.20	Apparent and Partial Molar Volumes at Infinite Dilution and Solid Liquid Equilibria of Dibenzothiophene + Alkane Systems
rhol	686.37	kg/m3	313.20	Apparent and Partial Molar Volumes at Infinite Dilution and Solid Liquid Equilibria of Dibenzothiophene + Alkane Systems
rhol	718.25	kg/m3	273.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	698.53	kg/m3	298.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	710.34	kg/m3	283.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa

rhol	706.35	kg/m3	288.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	702.35	kg/m3	293.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	698.34	kg/m3	298.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	694.30	kg/m3	303.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	690.24	kg/m3	308.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	686.17	kg/m3	313.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa

rhol	682.06	kg/m3	318.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	677.94	kg/m3	323.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	673.79	kg/m3	328.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	669.61	kg/m3	333.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	665.40	kg/m3	338.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	661.16	kg/m3	343.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa

rhol	656.88	kg/m3	348.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	652.57	kg/m3	353.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	648.23	kg/m3	358.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	643.83	kg/m3	363.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	698.76	kg/m3	298.15	Thermodynamics of Mixtures Containing Aromatic Alcohols. 1. Liquid Liquid Equilibria for (Phenylmethanol + Alkane) Systems
rhol	698.68	kg/m3	298.20	Isobaric Vapor Liquid Equilibrium of Binary Systems of Hexane or Octane with 1,2-Dimethylbenzene or 1,3-Dimethylbenzene at 101.3 kPa

rhol	702.54	kg/m3	293.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	698.49	kg/m3	298.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	694.44	kg/m3	303.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	690.37	kg/m3	308.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure
rhol	686.29	kg/m3	313.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure

rhol	682.19	kg/m3	318.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	678.08	kg/m3	323.15	Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, + 1-Propanol, + 1-Butanol, and + 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	
rhol	702.86	kg/m3	293.15	Excess Molar Volumes and Viscosities of Binary Systems of Butylcyclohexane with n-Alkanes (C7 to C14) at T = 293.15 K to 313.15 K	
rhol	698.85	kg/m3	298.15	Excess Molar Volumes and Viscosities of Binary Systems of Butylcyclohexane with n-Alkanes (C7 to C14) at T = 293.15 K to 313.15 K	
rhol	694.80	kg/m3	303.15	Excess Molar Volumes and Viscosities of Binary Systems of Butylcyclohexane with n-Alkanes (C7 to C14) at T = 293.15 K to 313.15 K	

rhol	690.73	kg/m3	308.15	Excess Molar Volumes and Viscosities of Binary Systems of Butylcyclohexane with n-Alkanes (C7 to C14) at T = 293.15 K to 313.15 K
rhol	686.63	kg/m3	313.15	Excess Molar Volumes and Viscosities of Binary Systems of Butylcyclohexane with n-Alkanes (C7 to C14) at T = 293.15 K to 313.15 K
rhol	698.64	kg/m3	298.15	Thermodynamics of Mixtures Containing a Strongly Polar Compound. 9. Liquid-Liquid Equilibria for epsilon-Caprolactam + Selected Alkanes
rhol	702.55	kg/m3	293.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	690.50	kg/m3	308.15	The density, the refractive index and the adjustment of the excess thermodynamic properties by means of the multiple linear regression method for the ternary system ethylbenzene-octane-propylbenzene

rhol	698.60	kg/m3	298.15	The density, the refractive index and the adjustment of the excess thermodynamic properties by means of the multiple linear regression method for the ternary system ethylbenzene-octane-propylbenzene
rhol	698.84	kg/m3	298.15	Liquid-liquid equilibria for (2-hydroxy benzaldehyde + n-alkane) mixtures. Intermolecular and proximity effects in systems containing hydroxyl and aldehyde groups
rhol	702.93	kg/m3	293.15	Surface tension and density of mixtures of m-xylene + n-alkane at 293.15 K: Analysis under the extended Langmuir and Shereshefsky models
rhol	698.40	kg/m3	298.15	Fully automatized apparatus for determining speed of sound for liquids in the temperature and pressure interval (283.15-343.15) K and (0.1-95) MPa
rhol	694.88	kg/m3	303.15	Thermodynamic behavior of thiophene with octane, 1-hexyl-3-methylimidazolium bromide, or 1-octyl-3-methylimidazolium bromide in dilute region at T = (288.15 to 303.15) K

rhol	698.92	kg/m3	298.15	Thermodynamic behavior of thiophene with octane, 1-hexyl-3-methylimidazolium bromide, or 1-octyl-3-methylimidazolium bromide in dilute region at T = (288.15 to 303.15) K
rhol	678.11	kg/m3	323.15	Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to 343.15) K at 0.1 MPa
rhol	703.00	kg/m3	293.15	Thermodynamic behavior of thiophene with octane, 1-hexyl-3-methylimidazolium bromide, or 1-octyl-3-methylimidazolium bromide in dilute region at T = (288.15 to 303.15) K
rhol	706.99	kg/m3	288.15	Thermodynamic behavior of thiophene with octane, 1-hexyl-3-methylimidazolium bromide, or 1-octyl-3-methylimidazolium bromide in dilute region at T = (288.15 to 303.15) K
rhol	698.86	kg/m3	298.15	Determination and correlation of (liquid + liquid) equilibria of ternary and quaternary systems with octane, decane, benzene and [BMpyr][DCA] at T = 298.15 K and atmospheric pressure
rhol	698.60	kg/m3	298.15	(Liquid + liquid) equilibrium at T = 298.15 K for ternary mixtures of alkane + aromatic compounds + imidazolium-based ionic liquids

rhol	698.55	kg/m3	298.15	Density, viscosity and excess molar volume of binary mixtures of tri-n-octylamine + diluents (n-heptane, n-octane, n-nonane, and n-decane) at various temperatures	
rhol	694.50	kg/m3	303.15	Density, viscosity and excess molar volume of binary mixtures of tri-n-octylamine + diluents (n-heptane, n-octane, n-nonane, and n-decane) at various temperatures	
rhol	702.60	kg/m3	293.15	Density, viscosity and excess molar volume of binary mixtures of tri-n-octylamine + diluents (n-heptane, n-octane, n-nonane, and n-decane) at various temperatures	
rhol	710.61	kg/m3	283.15	Density, viscosity and excess molar volume of binary mixtures of tri-n-octylamine + diluents (n-heptane, n-octane, n-nonane, and n-decane) at various temperatures	
rhol	698.60	kg/m3	298.15	Evaluation of ionic liquids as solvent for aromatic extraction: Experimental, correlation and COSMO-RS predictions	

rhol	698.75	kg/m3	298.15	Thermodynamics of mixtures containing amines. XI. Liquid + liquid equilibria and molar excess enthalpies at 298.15 K for N-methylaniline + hydrocarbon systems. Characterization in terms of DISQUAC and ERAS models
rhol	698.73	kg/m3	298.15	Excess molar enthalpies of the ternary mixtures: (tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether + n-octane) at the temperature 298.15 K
rhol	698.86	kg/m3	298.15	Application of the ionic liquid tributylmethylammonium bis(trifluoromethylsulfonyl)imide as solvent for the extraction of benzene from octane and decane at T = 298.15 K and atmospheric pressure
rhol	698.86	kg/m3	298.15	Study of the suitability of two ammonium-based ionic liquids for the extraction of benzene from its mixtures with aliphatic hydrocarbons.
rhol	686.39	kg/m3	313.15	Liquid liquid equilibria for the ternary system water + octane + 2-butyloxy-ethanol
rhol	702.64	kg/m3	293.15	Liquid liquid equilibria for the ternary system water + octane + 2-butyloxy-ethanol

rhol	698.61	kg/m3	298.15	Isobaric vapor-liquid equilibrium data of the binary systems of octane with p, o, m-xylene at 20 kPa	
rhol	698.51	kg/m3	298.10	Excess enthalpies of binary mixtures of 2-ethoxyethanol with four hydrocarbons at 298.15, 308.15, and 318.15K An experimental and theoretical study	
rhol	698.50	kg/m3	298.15	Densities and Kinematic Viscosities of a Quinary Regular Liquid System and Its Five Quaternary Subsystems at 293.15A K and 298.15A K	
rhol	702.50	kg/m3	293.15	Densities and Kinematic Viscosities of a Quinary Regular Liquid System and Its Five Quaternary Subsystems at 293.15A K and 298.15A K	
rhol	698.50	kg/m3	298.15	Densities and Kinematic Viscosities of One Quinary Regular Liquid System and Its Five Quaternary Sub-Systems at Temperatures (293.15 and 298.15) K	
rhol	702.50	kg/m3	293.15	Densities and Kinematic Viscosities of One Quinary Regular Liquid System and Its Five Quaternary Sub-Systems at Temperatures (293.15 and 298.15) K	
rhol	703.00	kg/m3	293.00	KDB	

rhol	698.64	kg/m3	298.15	Thermodynamics of mixtures containing amines VI. Liquid liquid equilibria for mixtures of o-toluidine + selected alkanes
rhol	714.31	kg/m3	278.15	Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa
rhol	682.20	kg/m3	318.15	The density, the refractive index and the adjustment of the excess thermodynamic properties by means of the multiple linear regression method for the ternary system ethylbenzene-octane-propylbenzene
sfust	93.19	J/mol×K	215.60	NIST Webbook
sfust	95.70	J/mol×K	215.80	NIST Webbook
sfust	95.85	J/mol×K	216.38	NIST Webbook
speedsl	1130.30	m/s	308.15	Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3)
speedsl	1151.00	m/s	303.15	Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3)
speedsl	1171.90	m/s	298.15	Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3)

speedsl	1193.00	m/s	293.15	Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3)
speedsl	1213.60	m/s	288.15	Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3)
srf	0.02	N/m	298.15	Interfacial tensions of binary mixtures of ethanol with octane, decane, dodecane, and tetradecane
srf	0.02	N/m	308.15	Interfacial tensions of binary mixtures of ethanol with octane, decane, dodecane, and tetradecane
srf	0.02	N/m	318.15	Interfacial tensions of binary mixtures of ethanol with octane, decane, dodecane, and tetradecane
srf	0.02	N/m	308.15	Surface Tension of Dialkyl Carbonates + (Alkanes or 1,4-Dimethylbenzene) and 1,4-Dimethylbenzene + Alkanes Binary Mixtures at T = 308.15 K
srf	0.02	N/m	293.20	KDB
srf	0.02	N/m	298.15	Analysis of Surface Tension, Density, and Speed of Sound for the Ternary Mixture Dimethyl Carbonate + p-Xylene + n-Octane

tcondl	0.12	W/m×K	334.80	Thermal Conductivities of [bmim][PF6], [hmim][PF6], and [omim][PF6] from 294 to 335 K at Pressures up to 20 MPa.
tcondl	0.12	W/m×K	314.70	Thermal Conductivities of [bmim][PF6], [hmim][PF6], and [omim][PF6] from 294 to 335 K at Pressures up to 20 MPa.
tcondl	0.13	W/m×K	294.50	Thermal Conductivities of [bmim][PF6], [hmim][PF6], and [omim][PF6] from 294 to 335 K at Pressures up to 20 MPa.

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	352.83	K	22.95	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	360.22	K	29.98	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures
tbp	366.89	K	37.87	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

tbp	371.89	K	44.82	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	376.93	K	52.81	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	380.87	K	59.85	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	384.36	K	66.66	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	388.38	K	75.23	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	391.59	K	82.72	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	
tbp	394.32	K	89.51	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures	

tbp	397.87	K	99.01	A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.44378e+01
Coeff. B	-3.45148e+03
Coeff. C	-4.72760e+01
Temperature range (K), min.	291.19
Temperature range (K), max.	425.47

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	6.57782e+01
Coeff. B	-6.98194e+03
Coeff. C	-7.37874e+00
Coeff. D	3.38092e-06
Temperature range (K), min.	216.38
Temperature range (K), max.	568.83

Datasets

Thermal conductivity, W/m/K

Temperature, K - Liquid	Pressure, kPa - Liquid	Thermal conductivity, W/m/K - Liquid
255.43	100.00	0.1386

255.34	5200.00	0.1409
255.48	10100.00	0.1427
255.80	20000.00	0.1458
255.12	30000.00	0.1482
274.91	100.00	0.1318
275.13	5000.00	0.1338
275.12	10200.00	0.1358
274.99	20200.00	0.1400
275.01	30200.00	0.1438
295.39	200.00	0.1265
295.22	5100.00	0.1287
295.33	10200.00	0.1309
295.18	20000.00	0.1351
295.13	30000.00	0.1391
315.13	100.00	0.1201
315.18	5000.00	0.1227
315.06	10000.00	0.1252
315.14	20200.00	0.1299
315.00	30000.00	0.1340
334.94	100.00	0.1143
334.97	5100.00	0.1170
335.05	10000.00	0.1194
334.94	20100.00	0.1245
334.94	30000.00	0.1288
355.08	100.00	0.1084
354.92	5000.00	0.1114
355.03	10000.00	0.1142
355.02	20000.00	0.1194
354.99	30000.00	0.1242
375.06	400.00	0.1019
374.87	5100.00	0.1052
375.08	10200.00	0.1082
374.90	20000.00	0.1147
375.26	30000.00	0.1198

Reference *https://www.doi.org/10.1021/acs.jced.9b00628*

Molar heat capacity at constant pressure, J/K/mol

Temperature, K - Liquid	Pressure, kPa - Liquid	Molar heat capacity at constant pressure, J/K/mol - Liquid
323.15	100.00	263.30

328.15	100.00	265.70
333.15	100.00	267.98
338.15	100.00	270.38
343.15	100.00	272.66
348.15	100.00	275.18
353.15	100.00	277.58
358.15	100.00	280.09
363.15	100.00	282.60
368.15	100.00	285.11
373.15	100.00	287.86
378.15	100.00	290.37
383.15	100.00	293.45
323.15	10130.00	260.90
328.15	10130.00	263.18
333.15	10130.00	265.58
338.15	10130.00	267.75
343.15	10130.00	270.04
348.15	10130.00	272.32
353.15	10130.00	274.49
358.15	10130.00	277.00
363.15	10130.00	279.17
368.15	10130.00	281.57
373.15	10130.00	283.97
378.15	10130.00	286.49
383.15	10130.00	288.88
388.15	10130.00	290.94
393.15	10130.00	292.88
398.15	10130.00	295.97

Reference *https://www.doi.org/10.1016/j.jct.2017.03.034*

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
273.15	100.00	0.0007050
273.15	10000.00	0.0007930
273.15	20000.00	0.0008710
273.15	30000.00	0.0009710
293.15	100.00	0.0005440
293.15	10000.00	0.0005960
293.15	20000.00	0.0006590

293.15	30000.00	0.0007170
313.15	100.00	0.0004290
313.15	10000.00	0.0004830
313.15	20000.00	0.0005360
313.15	30000.00	0.0005930
333.15	100.00	0.0003510
333.15	10000.00	0.0003880
333.15	20000.00	0.0004410
333.15	30000.00	0.0004820

Reference *https://www.doi.org/10.1007/s10765-006-0053-2*

Temperature, K	Pressure, kPa	Viscosity, Pa*s
303.15	101.33	0.0004918

Reference *https://www.doi.org/10.1016/j.fluid.2010.10.009*

Temperature, K	Pressure, kPa	Viscosity, Pa*s
298.24	100.00	0.0005080
308.72	5000.00	0.0004786
308.72	10000.00	0.0005043
323.93	100.00	0.0003829
348.64	100.00	0.0003034
349.46	5000.00	0.0003187
349.46	10000.00	0.0003364
394.81	5000.00	0.0002197
394.81	10000.00	0.0002334
436.86	5000.00	0.0001672
436.86	10000.00	0.0001799
467.46	5000.00	0.0001338
467.46	10000.00	0.0001458

Reference *https://www.doi.org/10.1016/j.jct.2006.01.012*

Temperature, K	Pressure, kPa	Viscosity, Pa*s
323.15	100.00	0.0003892
323.15	10096.00	0.0004400
323.15	20701.00	0.0004950
323.15	30654.00	0.0005300
323.15	41050.00	0.0005760
323.15	54543.00	0.0006330

323.15	68036.00	0.0006930
348.14	100.00	0.0003011
348.14	7551.00	0.0003338
348.14	13680.00	0.0003513
348.14	20500.00	0.0003780
348.14	27466.00	0.0004008
348.14	27469.00	0.0004000
348.14	34305.00	0.0004240
348.14	34314.00	0.0004240
348.14	41226.00	0.0004490
348.14	41227.00	0.0004490
348.14	48070.00	0.0004710
348.14	55000.00	0.0004980
348.14	55004.00	0.0004990
348.14	61992.00	0.0005220
348.14	61994.00	0.0005330
373.12	6723.00	0.0002757
373.12	6724.00	0.0002758
373.12	20592.00	0.0003205
373.12	20599.00	0.0003217
373.12	27397.00	0.0003429
373.12	27402.00	0.0003429
373.12	34317.00	0.0003646
373.12	34324.00	0.0003649
373.12	41184.00	0.0003810
373.12	41207.00	0.0003800
373.12	41221.00	0.0003840
373.12	48024.00	0.0004090
373.12	48032.00	0.0004090
373.12	54853.00	0.0004330
373.12	54866.00	0.0004320
373.12	61686.00	0.0004470
373.12	61704.00	0.0004490
423.09	14725.00	0.0002290
423.09	27159.00	0.0002520
423.09	40831.00	0.0002870
423.09	54591.00	0.0003440
423.09	68420.00	0.0003520

Reference *https://www.doi.org/10.1021/je0503296*

Temperature, K	Pressure, kPa	Viscosity, Pa*s
298.15	100.00	0.0005100

298.15	20900.00	0.0006310
298.15	40600.00	0.0007530
298.15	60900.00	0.0008890
298.15	80000.00	0.0010400
298.15	80100.00	0.0010300
298.15	100300.00	0.0011900
298.15	120100.00	0.0013600
298.15	139700.00	0.0015550
298.15	139900.00	0.0015500
298.15	159500.00	0.0017500
298.15	179600.00	0.0019800
298.15	197200.00	0.0022100
323.15	100.00	0.0003870
323.15	20100.00	0.0004770
323.15	40000.00	0.0005700
323.15	60600.00	0.0006720
323.15	80100.00	0.0007780
323.15	101000.00	0.0008970
323.15	120300.00	0.0010100
323.15	141500.00	0.0011500
323.15	141600.00	0.0011500
323.15	160300.00	0.0012800
323.15	180700.00	0.0014400
323.15	201800.00	0.0016100
323.15	100.00	0.0003870
323.15	80100.00	0.0007810
348.15	100.00	0.0003040
348.15	20400.00	0.0003810
348.15	40800.00	0.0004570
348.15	60700.00	0.0005340
348.15	80100.00	0.0006140
348.15	101000.00	0.0007060
348.15	120400.00	0.0007940
348.15	140200.00	0.0008930
348.15	140300.00	0.0008920
348.15	160700.00	0.0009970
348.15	180800.00	0.0011100
348.15	201900.00	0.0012300
348.15	100.00	0.0003040
348.15	80100.00	0.0006140
373.15	100.00	0.0002450
373.15	20600.00	0.0003120
373.15	39900.00	0.0003730
373.15	60000.00	0.0004370

373.15	80200.00	0.0005030
373.15	101100.00	0.0005760
373.15	120000.00	0.0006450
373.15	140900.00	0.0007260
373.15	160300.00	0.0008040
373.15	180500.00	0.0008900
373.15	200900.00	0.0009810
398.15	21600.00	0.0002630
398.15	42000.00	0.0003200
398.15	79400.00	0.0004260
398.15	80900.00	0.0004300
398.15	120900.00	0.0005500
398.15	159500.00	0.0006740
398.15	160200.00	0.0006760
398.15	192300.00	0.0007860
398.15	21600.00	0.0002650
423.15	21400.00	0.0002260
423.15	41100.00	0.0002760
423.15	80400.00	0.0003720
423.15	80700.00	0.0003720
423.15	121500.00	0.0004760
423.15	159700.00	0.0005810
423.15	161300.00	0.0005870
423.15	192500.00	0.0006760
423.15	21400.00	0.0002260
448.15	21000.00	0.0001920
448.15	41200.00	0.0002390
448.15	80500.00	0.0003240
448.15	81400.00	0.0003260
448.15	121500.00	0.0004140
448.15	159400.00	0.0005040
448.15	160300.00	0.0005040
448.15	193000.00	0.0005850
448.15	21000.00	0.0001920
473.15	19300.00	0.0001630
473.15	20000.00	0.0001640
473.15	41100.00	0.0002100
473.15	81100.00	0.0002890
473.15	81500.00	0.0002890
473.15	121100.00	0.0003670
473.15	160700.00	0.0004480
473.15	161300.00	0.0004490
473.15	193400.00	0.0005160

Reference *https://www.doi.org/10.1021/je800417q*

Mass density, kg/m3

Temperature, K - Liquid	Pressure, kPa - Liquid	Mass density, kg/m3 - Liquid
288.15	100.00	706.6
288.15	5000.00	710.9
288.15	10000.00	714.7
288.15	20000.00	721.7
288.15	30000.00	728.1
288.15	40000.00	734.3
298.15	100.00	698.6
298.15	5000.00	703.4
298.15	10000.00	707.0
298.15	20000.00	714.6
298.15	30000.00	721.4
298.15	40000.00	727.7
308.15	100.00	690.5
308.15	5000.00	695.0
308.15	10000.00	699.4
308.15	20000.00	707.3
308.15	30000.00	714.5
308.15	40000.00	720.8

Reference *https://www.doi.org/10.1016/j.jct.2012.11.011*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.15	101.00	702.58
303.15	101.00	694.48
313.15	101.00	686.32
323.15	101.00	678.09
288.15	100.00	706.5
288.15	1000.00	707.2
288.15	5000.00	710.3
288.15	10000.00	714.1
288.15	15000.00	717.7
288.15	20000.00	721.2
288.15	25000.00	724.5
288.15	30000.00	727.7

288.15	35000.00	730.8
288.15	40000.00	733.7
288.15	45000.00	736.5
288.15	50000.00	739.2
288.15	55000.00	741.7
288.15	60000.00	744.1
293.15	100.00	702.5
293.15	1000.00	703.2
293.15	5000.00	706.4
293.15	10000.00	710.3
293.15	15000.00	714.0
293.15	20000.00	717.6
293.15	25000.00	721.0
293.15	30000.00	724.3
293.15	35000.00	727.5
293.15	40000.00	730.5
293.15	45000.00	733.3
293.15	50000.00	736.0
293.15	55000.00	738.6
293.15	60000.00	741.0
298.15	100.00	698.5
298.15	1000.00	699.2
298.15	5000.00	702.5
298.15	10000.00	706.5
298.15	15000.00	710.3
298.15	20000.00	714.0
298.15	25000.00	717.6
298.15	30000.00	720.9
298.15	35000.00	724.1
298.15	40000.00	727.2
298.15	45000.00	730.1
298.15	50000.00	732.9
298.15	55000.00	735.5
298.15	60000.00	738.0
303.15	100.00	694.4
303.15	1000.00	695.2
303.15	5000.00	698.6
303.15	10000.00	702.7
303.15	15000.00	706.7
303.15	20000.00	710.5
303.15	25000.00	714.1
303.15	30000.00	717.5
303.15	35000.00	720.9
303.15	40000.00	724.0

303.15	45000.00	727.0
303.15	50000.00	729.8
303.15	55000.00	732.5
303.15	60000.00	735.0
308.15	100.00	690.4
308.15	1000.00	691.2
308.15	5000.00	694.7
308.15	10000.00	698.9
308.15	15000.00	703.0
308.15	20000.00	706.9
308.15	25000.00	710.6
308.15	30000.00	714.2
308.15	35000.00	717.5
308.15	40000.00	720.8
308.15	45000.00	723.8
308.15	50000.00	726.7
308.15	55000.00	729.4
308.15	60000.00	731.9
313.15	100.00	686.3
313.15	1000.00	687.1
313.15	5000.00	690.8
313.15	10000.00	695.1
313.15	15000.00	699.3
313.15	20000.00	703.3
313.15	25000.00	707.2
313.15	30000.00	710.8
313.15	35000.00	714.3
313.15	40000.00	717.6
313.15	45000.00	720.7
313.15	50000.00	723.6
313.15	55000.00	726.3
313.15	60000.00	728.9
318.15	100.00	682.2
318.15	1000.00	683.0
318.15	5000.00	686.8
318.15	10000.00	691.3
318.15	15000.00	695.6
318.15	20000.00	699.7
318.15	25000.00	703.6
318.15	30000.00	707.4
318.15	35000.00	711.0
318.15	40000.00	714.3
318.15	45000.00	717.5
318.15	50000.00	720.5

318.15	55000.00	723.3
318.15	60000.00	725.9
323.15	100.00	678.0
323.15	1000.00	678.9
323.15	5000.00	682.8
323.15	10000.00	687.4
323.15	15000.00	691.9
323.15	20000.00	696.1
323.15	25000.00	700.2
323.15	30000.00	704.0
323.15	35000.00	707.7
323.15	40000.00	711.1
323.15	45000.00	714.4
323.15	50000.00	717.4
323.15	55000.00	720.3
323.15	60000.00	722.9
328.15	100.00	673.8
328.15	1000.00	674.7
328.15	5000.00	678.7
328.15	10000.00	683.5
328.15	15000.00	688.1
328.15	20000.00	692.5
328.15	25000.00	696.7
328.15	30000.00	700.7
328.15	35000.00	704.4
328.15	40000.00	707.9
328.15	45000.00	711.3
328.15	50000.00	714.4
328.15	55000.00	717.3
328.15	60000.00	719.9
333.15	100.00	669.7
333.15	1000.00	670.6
333.15	5000.00	674.8
333.15	10000.00	679.7
333.15	15000.00	684.5
333.15	20000.00	689.0
333.15	25000.00	693.2
333.15	30000.00	697.3
333.15	35000.00	701.1
333.15	40000.00	704.7
333.15	45000.00	708.1
333.15	50000.00	711.3
333.15	55000.00	714.2
333.15	60000.00	716.9

343.15	100.00	661.2
343.15	1000.00	662.2
343.15	5000.00	666.6
343.15	10000.00	671.9
343.15	15000.00	677.0
343.15	20000.00	681.8
343.15	25000.00	686.3
343.15	30000.00	690.6
343.15	35000.00	694.6
343.15	40000.00	698.4
343.15	45000.00	701.9
343.15	50000.00	705.2
343.15	55000.00	708.2
343.15	60000.00	711.0
353.15	100.00	652.6
353.15	1000.00	653.7
353.15	5000.00	658.4
353.15	10000.00	664.1
353.15	15000.00	669.4
353.15	20000.00	674.5
353.15	25000.00	679.3
353.15	30000.00	683.9
353.15	35000.00	688.1
353.15	40000.00	692.1
353.15	45000.00	695.7
353.15	50000.00	699.1
353.15	55000.00	702.3
353.15	60000.00	705.1
363.15	100.00	643.8
363.15	1000.00	645.0
363.15	5000.00	650.1
363.15	10000.00	656.1
363.15	15000.00	661.9
363.15	20000.00	667.3
363.15	25000.00	672.4
363.15	30000.00	677.2
363.15	35000.00	681.6
363.15	40000.00	685.8
363.15	45000.00	689.6
363.15	50000.00	693.1
363.15	55000.00	696.3
363.15	60000.00	699.2
373.15	100.00	634.7
373.15	1000.00	636.3

373.15	5000.00	641.8
373.15	10000.00	648.2
373.15	15000.00	654.4
373.15	20000.00	660.1
373.15	25000.00	665.5
373.15	30000.00	670.6
373.15	35000.00	675.3
373.15	40000.00	679.6
373.15	45000.00	683.6
373.15	50000.00	687.3
373.15	55000.00	690.5
373.15	60000.00	693.5
393.15	100.00	616.3
393.15	1000.00	618.1
393.15	5000.00	624.3
393.15	10000.00	631.8
393.15	15000.00	638.8
393.15	20000.00	645.3
393.15	25000.00	651.5
393.15	30000.00	657.1
393.15	35000.00	662.3
393.15	40000.00	667.1
393.15	45000.00	671.4
393.15	50000.00	675.3
393.15	55000.00	678.8
393.15	60000.00	681.8
413.15	1000.00	599.3
413.15	5000.00	606.6
413.15	10000.00	615.1
413.15	15000.00	623.1
413.15	20000.00	630.6
413.15	25000.00	637.4
413.15	30000.00	643.8
413.15	35000.00	649.6
413.15	40000.00	654.8
413.15	45000.00	659.5
413.15	50000.00	663.6
413.15	55000.00	667.2
413.15	60000.00	670.2

Reference *https://www.doi.org/10.1016/j.jct.2016.04.014*

Pressure, kPa	**Temperature, K**	**Mass density, kg/m3**

100.00	298.15	698.71

Reference *https://www.doi.org/10.1016/j.jct.2017.09.027*

Temperature, K	Pressure, kPa	Mass density, kg/m3
293.21	1996.00	703.76
293.22	3994.00	705.71
293.22	6001.00	707.57
293.21	8004.00	709.25
293.22	10009.00	710.88
293.21	13010.00	713.21
293.22	16009.00	715.43
293.22	19011.00	717.57
293.22	22012.00	719.76
293.22	25004.00	721.88
293.22	28004.00	724.19
293.21	30002.00	725.54
313.11	2007.00	688.27
313.11	4000.00	690.26
313.10	5999.00	692.07
313.11	7997.00	693.85
313.11	10009.00	695.59
313.11	13005.00	698.26
313.11	16009.00	700.84
313.10	19006.00	703.5
313.10	21998.00	706.05
313.10	25000.00	708.39
313.10	27998.00	710.67
313.10	30001.00	712.15
333.01	2005.00	672.19
333.01	3993.00	674.49
333.01	5998.00	676.69
333.01	8014.00	678.82
333.01	10007.00	680.8
333.01	13004.00	683.6
333.01	16002.00	686.47
333.01	19006.00	689.24
333.00	22003.00	691.91
333.01	25006.00	694.49
333.01	28003.00	696.99
333.01	30012.00	698.61
353.12	2008.00	655.38
353.12	3999.00	657.99

353.12	6000.00	660.5
353.12	8001.00	663.11
353.12	10006.00	665.45
353.13	13003.00	668.83
353.13	16004.00	672.05
353.12	19000.00	674.99
353.12	22003.00	677.96
353.12	25005.00	680.81
353.12	28008.00	683.56
353.12	30004.00	685.34

Reference *https://www.doi.org/10.1021/acs.jced.7b00650*

Speed of sound, m/s

Temperature, K - Liquid	Pressure, kPa - Liquid	Speed of sound, m/s - Liquid
296.97	100.00	1173.0
303.61	100.00	1151.7
318.37	100.00	1089.5
332.50	100.00	1029.1
347.68	100.00	970.1
362.53	100.00	912.3
377.50	100.00	855.4
391.61	100.00	806.4
297.28	4000.00	1196.5
303.53	4000.00	1172.7
318.30	4000.00	1121.2
332.26	4000.00	1066.7
347.40	4000.00	1007.2
362.28	4000.00	955.4
377.03	4000.00	909.7
390.80	4000.00	860.1
407.38	4000.00	802.1
412.59	4000.00	786.2
422.73	4000.00	747.2
431.82	4000.00	720.1
442.23	4000.00	682.1
452.55	4000.00	646.4
462.37	4000.00	614.7
472.05	4000.00	579.7
482.98	4000.00	543.0

493.41	4000.00	503.5
502.54	4000.00	477.8
513.00	4000.00	439.1
522.92	4000.00	407.7
532.23	4000.00	373.2
537.40	4000.00	354.7
542.87	4000.00	337.8
546.68	4000.00	325.7
550.42	4000.00	315.2
557.27	4000.00	290.5
561.20	4000.00	276.7
567.01	4000.00	257.3
570.16	4000.00	245.6
297.34	7000.00	1222.0
303.33	7000.00	1199.7
318.35	7000.00	1141.4
332.33	7000.00	1089.1
347.29	7000.00	1031.9
362.17	7000.00	979.6
377.33	7000.00	933.5
391.00	7000.00	892.3
407.46	7000.00	836.9
412.65	7000.00	818.8
422.88	7000.00	784.5
431.88	7000.00	754.2
442.49	7000.00	723.8
452.64	7000.00	689.9
462.45	7000.00	660.1
472.15	7000.00	631.6
483.04	7000.00	599.7
493.48	7000.00	565.4
502.73	7000.00	537.1
513.26	7000.00	506.5
523.46	7000.00	478.6
532.36	7000.00	450.6
537.75	7000.00	436.8
543.83	7000.00	422.0
547.82	7000.00	406.8
552.01	7000.00	398.1
558.19	7000.00	380.2
561.87	7000.00	372.6
566.82	7000.00	356.5
570.20	7000.00	347.6
573.03	7000.00	341.0

575.57	7000.00	332.6
578.92	7000.00	323.1
297.38	10000.00	1242.7
303.38	10000.00	1213.7
318.11	10000.00	1164.8
332.32	10000.00	1108.6
347.37	10000.00	1053.4
362.12	10000.00	1006.2
377.35	10000.00	960.5
391.09	10000.00	917.4
407.45	10000.00	865.2
412.71	10000.00	846.7
422.61	10000.00	818.2
431.92	10000.00	789.8
442.65	10000.00	755.7
452.77	10000.00	728.1
465.54	10000.00	695.8
472.28	10000.00	676.9
483.12	10000.00	646.1
493.62	10000.00	614.7
502.95	10000.00	590.5
513.46	10000.00	563.1
523.60	10000.00	533.3
532.59	10000.00	511.4
538.10	10000.00	497.5
544.52	10000.00	483.1
548.60	10000.00	471.5
552.52	10000.00	465.1
558.64	10000.00	448.0
562.27	10000.00	441.3
566.96	10000.00	430.4
570.48	10000.00	424.0
573.37	10000.00	414.0
576.28	10000.00	408.7
579.36	10000.00	400.8
297.35	12000.00	1250.8
303.42	12000.00	1233.0
318.37	12000.00	1184.5
332.56	12000.00	1128.3
347.46	12000.00	1072.7
363.67	12000.00	1027.8
379.75	12000.00	973.9
394.67	12000.00	920.8
403.71	12000.00	896.1

414.33	12000.00	864.0
422.69	12000.00	844.3
431.77	12000.00	815.2
442.73	12000.00	779.7
452.88	12000.00	750.8
462.61	12000.00	724.2
472.38	12000.00	693.8
483.14	12000.00	672.7
493.73	12000.00	645.0
503.17	12000.00	616.5
513.67	12000.00	591.5
523.17	12000.00	570.6
532.79	12000.00	545.1
538.41	12000.00	530.4
543.27	12000.00	520.1
549.27	12000.00	505.4
552.88	12000.00	500.8
559.37	12000.00	486.2
562.71	12000.00	477.8
567.10	12000.00	471.0
570.75	12000.00	461.0
573.60	12000.00	454.4
576.53	12000.00	451.3
579.57	12000.00	444.1

Reference *https://www.doi.org/10.1021/acs.jced.7b00712*

Sources

Measurement of Critical Properties for Binary and Ternary Mixtures Containing n-Butanol and n-Alkane: https://www.doi.org/10.1021/acs.jced.8b00585

Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide: https://www.doi.org/10.1016/j.jct.2013.07.004

Determination of Infinite Dilution Activity Coefficients of Several Organic Solutes in Ionic Liquid 1-Octyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide by Inverse Gas Chromatography: https://www.doi.org/10.1021/acs.jced.7b00244

Selective separation of aliphatics from aromatics using supported ionic liquid membranes: https://www.doi.org/10.1016/j.jct.2015.12.033

Measurement of activity coefficients at infinite dilution for organic solutes, water and alcohols in the ionic liquid 1-ethyl-3-methylpyridinium bis(trifluoromethylsulfonyl)imide: 15 K to the atmospheric pressure: https://www.doi.org/10.1016/j.jct.2014.04.024

Separation of toluene from alkanes using 1-ethyl-3-methylpyridinium ethylsulfate and 1-ethyl-3-methylpyridinium dicyanamide ionic liquids at T = 298.15 K and atmospheric pressure: https://www.doi.org/10.1016/j.jct.2010.01.008

Determination of activity coefficients at infinite dilution of organic solutes in the ionic liquid tributylmethylphosphonium methylsulfate by gas–liquid chromatography: https://www.doi.org/10.1016/j.fluid.2006.03.002

Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-3-methylpyridinium bis(trifluoromethylsulfonyl)imide using gas–liquid chromatography: https://www.doi.org/10.1016/j.jct.2013.10.026

Activity Coefficients at Infinite Dilution for Hydrocarbons in Isomeric Alcohols https://www.doi.org/10.1021/je1005517

Measurement of activity coefficients at infinite dilution of organic solutes in 2-(2-methoxyethyl)-1-ethyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide using gas–liquid chromatography: https://www.doi.org/10.1016/j.jct.2018.07.024

Ionic liquids for separation of binary hydrocarbons mixtures based on gamma infinity data measurements: https://www.doi.org/10.1016/j.jct.2005.04.010

Measurement of VLE and c1 data and prediction of their thermodynamic behavior using original UNIFAC, mod. UNIFAC(Do) and COSMO-RS(Ol): https://www.doi.org/10.1016/j.jct.2011.11.021

Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:

Solubility Measurement and Correlation for epsilon-2,4,6,8,10,12-Hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Different Alkanes/Aromatic Hydrocarbon + Ethyl Acetate Binary Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00353
Interfacial Tension between Methane and Octane at Elevated Pressure at Five Temperatures from (274.2 to 422.2) K:	https://www.doi.org/10.1021/je200560w
Activity coefficients at infinite dilution of organic solutes in N-alkylpyridinium bis(trifluoromethylsulfonyl)imide ([CnPY][NTf2], n = 2, 4, 5) using gas liquid chromatography:	https://www.doi.org/10.1016/j.jct.2010.06.009
Measuring Activity Coefficient at Infinite Dilution of Hydrocarbons in Ionic Liquids and Evaluation of Other Thermodynamic Properties using Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.9b00170
(Liquid + liquid) equilibrium at T = 298.15 K for ternary mixtures of alkane + aromatic compounds + imidazolium-based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.01.022
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in 1-Ethyl-3-methylimidazolium Diethylphosphate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je200266f
Prediction and measurement of phase equilibria for the extraction of BTX from naphtha reformate using BMIMPF6 ionic liquid:	https://www.doi.org/10.1016/j.fluid.2013.11.034
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in 1-(2-Hydroxyethyl)-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je9002724
Ternary Liquid-Liquid Equilibria for Mixtures of an Aromatic + an Aliphatic Hydrocarbon + 4-Methyl-N-butylpyridinium Tetrafluoroborate:	https://www.doi.org/10.1021/je0601980
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium methyl sulfate:	https://www.doi.org/10.1016/j.jct.2014.04.020
Experimental Vapor-Liquid Equilibria for the Carbon Dioxide + Octane and Carbon Dioxide + Decane Systems:	https://www.doi.org/10.1021/je060111z
Experimental and theoretically study of interaction between organic compounds and tricyanomethanide based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.12.027
Determination of Activity Coefficients at Infinite Dilution of Solutes in the Ionic Liquid 1-Butyl-3-methylimidazolium Octyl Sulfate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je0500375
Application of the ionic liquid tributylmethylammonium bis(trifluoromethylsulfonyl)imide as solvent for the extraction of benzene from [illegible] at 298.15 K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2016.02.028
Thermodynamic study of (perfluoroalkane + alkane) mixtures: [illegible]	https://www.doi.org/10.1016/j.jct.2007.04.002
Determination of Activity Coefficients at Infinite Dilution of Polar and Nonpolar Solutes in the Ionic Liquid 1-Ethyl-3-methyl-imidazolium Bis(trifluoromethylsulfonyl) imidate Using Gas-Liquid Chromatography at the Temperature 303.15 K or 318.15 K:	https://www.doi.org/10.1021/je0498107
Separation of binary mixtures based on gamma infinity data using [illegible] ionic liquid and thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2017.12.012
Liquid-liquid equilibria of mixtures with ethyl lactate and various hydrocarbons:	https://www.doi.org/10.1016/j.fluid.2012.02.004
Anomalous nucleation near a liquid-liquid phase boundary created by a rapid heat pulse:	https://www.doi.org/10.1016/j.fluid.2015.12.046
Isobaric vapor-liquid equilibria for the n-heptane + ethylene glycol monopropyl ether and n-octane + ethylene glycol monopropyl ether systems:	https://www.doi.org/10.1016/j.fluid.2010.08.002
Experimental densities, refractive indices, and speeds of sound of 12 binary mixtures containing alkanes and aromatic compounds at T = 313.15 K:	https://www.doi.org/10.1016/j.jct.2009.03.009
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2017.03.004
Measurement and Correlation of Mutual Solubilities for High Viscosity Binary Systems: Aniline + Methylcyclohexane, Phenol + Heptane, Phenol + Octane, and Glycerol + 1-Pentanol:	https://www.doi.org/10.1021/je034043h
Infinite Dilution Activity Coefficients for Trihexyltetradecyl Phosphonium Ionic Liquids: Measurements and COSMO-RS Prediction:	https://www.doi.org/10.1021/je0602925
Measurement of activity coefficients at infinite dilution using polar and non-polar solutes in the ionic liquid 1-methyl-3-octyl-imidazolium diethyleneglycolmonomethylethersulfate at T = (288.15, 298.15, and 313.15) K:	https://www.doi.org/10.1016/j.jct.2005.07.003
Thermodynamics of Mixtures Containing a Strongly Polar Compound. [illegible] Liquid-Liquid Equilibria for [illegible] and Selected Alkanes:	https://www.doi.org/10.1021/je900785z
Thermophysical Properties of Binary Mixtures of 2-Methyl-1-propanol with Hexane, Octane, and Decane at 298.15 K:	https://www.doi.org/10.1021/je060389r
Gas-liquid chromatography measurements of activity coefficients at infinite dilution of hydrocarbons and alkanols in 1-alkyl-3-methylimidazolium bis(oxalato)borate:	https://www.doi.org/10.1016/j.fluid.2010.08.016
A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Activity Coefficients at Infinite Dilution of Organic Solutes and Water in Tributylethylphosphonium Diethylphosphate Using Gas Liquid Chromatography: Thermodynamic Properties of Mixtures Containing Ionic Liquids:	https://www.doi.org/10.1021/acs.jced.5b00980
Phase Equilibria in Systems of Morpholine, Acetonitrile, and n-Alkanes:	https://www.doi.org/10.1021/acs.jced.5b00175
Thermodynamic study of molecular interaction-selectivity in separation processes based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2018.02.014
Densities and Viscosities of Binary Mixtures of JP-10 with n-Octane or n-Decane at Several Temperatures:	https://www.doi.org/10.1021/je800348s
New ionic liquid [P4,4,4,4][NTf2] in bio-butanol extraction on investigation of limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.07.028
Surface Tension of Dialkyl Carbonates + (Alkanes or 1,4-Dimethylbenzene) and 1,4-Dimethylbenzene + Alkanes Binary Mixtures at T = 308.15 K:	https://www.doi.org/10.1021/je301282p
Ternary (liquid + liquid) equilibria for mixtures of (methanol + aniline + n-octane or n-dodecane) at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2007.05.018
Solubilities of 1-Hexyl-3-methylimidazole Nitrate and 1-Octyl-3-methylimidazole Nitrate in Selected Solvents:	https://www.doi.org/10.1021/je200151j
Gas-liquid chromatography measurements of activity coefficients at infinite dilution of various organic solutes and water in tri-iso-butylmethylphosphonium tosylate ionic liquid:	https://www.doi.org/10.1016/j.jct.2010.01.004
Vapour pressures and excess functions of (3,5; 2,6)dimethylpyridine + n-hexane, n-heptane and n-octane: measurement and prediction:	https://www.doi.org/10.1016/j.fluid.2005.03.028
Measurements of activity coefficients at infinite dilution for organic solutes in two quaternary ammonium-based ionic liquids [DDA][ClO4] and [DDA][BF4]:	https://www.doi.org/10.1016/j.fluid.2018.11.011
Excess volumes, densities, speeds of sound and viscosities for the binary systems of diisopropyl ether with hydrocarbons at 303.15K:	https://www.doi.org/10.1016/j.tca.2005.10.008
Experimental and predicted excess molar enthalpies for 1,4-dioxane + octane + cyclohexane at 303.15 K:	https://www.doi.org/10.1016/j.fluid.2006.03.020
Speed of sound and thermodynamic properties of ternary liquid mixture octane + dodecane + hexadecane:	https://www.doi.org/10.1016/j.fluid.2015.07.008

A novel dynamic recirculating apparatus for vapour-liquid equilibrium measurements at moderate pressures and temperatures:	https://www.doi.org/10.1016/j.fluid.2013.07.044
Isothermal Bubble Pressure Data for the Binary System of C2F6 and n-Octane:	https://www.doi.org/10.1021/acs.jced.8b00825
Isobaric Vapor-Liquid Equilibria of Hexane + 1-Decene and Octane + 1-Decene Mixtures:	https://www.doi.org/10.1021/je700445y
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylmorpholinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.08.030
Thermodynamics of Mixtures Containing Ethers. Part III. Liquid-Liquid Equilibria for 2,5,8,11-Tetraoxadodecane or 2,5,8,11,14-Pentaoxapentadecane + Selected N-Alkanes:	https://www.doi.org/10.1021/je049903o
Separation of aromatic hydrocarbons from alkanes using ammonium ionic liquid C2NTf2 at T = 298.15 K:	https://www.doi.org/10.1016/j.fluid.2007.06.025
A Digital Variable-Angle Rolling-Ball Viscometer for Measurement of Viscosity, Density, and Bubble-Point Pressure of CO2 and Organic Liquid Mixtures:	https://www.doi.org/10.1007/s10765-008-0542-6
Separation of benzene from alkanes using 1-ethyl-3-methylpyridinium ethylsulfate ionic liquid at several temperatures and atmospheric pressure: Effect of the size of the aliphatic hydrocarbons:	https://www.doi.org/10.1016/j.jct.2009.07.017
Measurements of the Liquid Viscosities of Mixtures of n-Butane, n-Hexane, and n-Octane with Squalane to 30 MPa:	https://www.doi.org/10.1007/s10765-006-0053-2
Excess Enthalpy, Density, and Speed of Sound for the Ternary Mixture Methyl tert-Butyl Ether (1) + Butan-1-ol (2) + Octane (3):	https://www.doi.org/10.1021/je8004613
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Ethyl-3-methylimidazolium Tris(pentafluoroethyl)trifluorophosphate [EMIM][FAP] Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je9008443
Measurement and Prediction of Vapor-Liquid Equilibria in Ternary Systems Containing [illegible] Cyclohexylamine, and Cyclohexanol at 20 kPa:	https://www.doi.org/10.1021/acs.jced.7b00721
Isobaric vapor-liquid equilibrium data of the binary systems of octane with p, [illegible]	https://www.doi.org/10.1016/j.fluid.2013.01.015
Determination of Activity Coefficients at Infinite Dilution of Solutes in the Ionic Liquid, Trihexyltetradecylphosphonium Bis(trifluoromethylsulfonyl) Imide, Using Gas-Liquid Chromatography at T = (303.15, 308.15, 313.15, and 318.15) K:	https://www.doi.org/10.1021/je800105r
Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl [illegible]	https://www.doi.org/10.1021/je100832h
Liquid-Liquid Equilibria for Monoethylene Glycol + Water + Alkane Systems in the Range (278 to 313) K and Atmospheric Pressure:	https://www.doi.org/10.1021/je034270j
The Yaws Handbook of Vapor Pressure	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
Evaluation of the Performance of Trigeminal Tricationic Ionic Liquids for Separation Problems:	https://www.doi.org/10.1021/je201129y
Excess Enthalpies of the Ternary Mixtures: Diisopropyl Ether + 3-Methylpentane + (Octane or Decane) at 298.15 K:	https://www.doi.org/10.1021/je030161t
Activity coefficients at infinite dilution of organic solutes in 1-octyl-3-methylimidazolium nitrate using gas-liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2012.05.006
Bubble Point Measurements of n-Butane + n-Octane and n-Butane + n-Nonane Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.5b00308
(Liquid + liquid) equilibria for ternary mixtures of (ethylene glycol + toluene + n-octane):	https://www.doi.org/10.1016/j.jct.2008.03.014
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-hydroxyethyl)-3-methylimidazolium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2013.05.008
Densities and Excess Molar Properties of Dimethyl Carbonate with Alkanes (C6 to C10) and VLE of Dimethyl Carbonate with Alkanes (C9 to C10) at 101.3 kPa:	https://www.doi.org/10.1021/je0301794
Volumetric Properties for the Binary Systems Hexane + Octane and Hexane [illegible] at High Temperatures and Pressures:	https://www.doi.org/10.1021/je200721d
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 5. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in 1-Methyl-3-butyl-imidazolium Bis(trifluoromethylsulfonyl) Imide Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je050125p
THERMODYNAMICS OF MIXTURES CONTAINING ALKOXYETHANOLS. XXVIII. LIQUID-LIQUID EQUILIBRIA FOR 2-PHENOXYETHANOL + SELECTED ALKANES:	https://www.doi.org/10.1016/j.tca.2011.04.012
tetramethylammonium chloride + glycerol deep eutectic solvent as separation agent for organic mixtures: Assessment from experimental limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.06.003
Activity Coefficients at Infinite Dilution of Organic Solutes in [illegible]-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je100952p
[illegible] 1-Alkyl-3-methylimidazolium Hexafluorophosphate in Hydrocarbons:	https://www.doi.org/10.1021/je020145g
Bubble Pressure Measurement and Prediction for n-Hexadecane and n-Eicosane + Cyclohexane, Methylcyclohexane, and Ethylcyclohexane Binary Mixtures from 303.15 to 393.15 K:	https://www.doi.org/10.1021/acs.jced.8b00125
Infinite Dilution Activity Coefficients of Hydrocarbons in Triethylene Glycol and Tetraethylene Glycol:	https://www.doi.org/10.1021/je020196d
Excess Molar Enthalpies for Binary Mixtures of Ethanol + Acetone, + Octane, + Cyclohexane and 1-Propanol + Acetone, + Octane, + Heptane at 323.15:	https://www.doi.org/10.1021/je020149l
Activity coefficients at infinite dilution of organic solutes in the ionic liquid trihexyltetradecylphosphonium hexafluorophosphate using gas-liquid chromatography at T = (313.15, 333.15, 353.15 and 373.15) K:	https://www.doi.org/10.1016/j.jct.2012.01.004
Measurement of Activity Coefficient at Infinite Dilution of Hydrocarbons in Sulfolane Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je700139p
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-hexyl-2,3-dimethylimidazolium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2016.07.021
Experimental Liquid Densities of n-Pentane, n-Octane, and n-Nonane and Their Binary Mixtures from (273.15 to 363.15) K at 0.1 MPa:	https://www.doi.org/10.1021/je200373q
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
(Liquid + liquid) equilibria for (water + 1-propanol + diisopropyl ether + octane, or methylbenzene, or heptane) systems at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2015.03.003
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.08.016
Physical properties of 3-methyl-N-butylpyridinium tetracyanoborate and ternary LLE data of [3-mebupy]B(CN)4 with an aromatic and an aliphatic hydrocarbon at T = 303.2 K and 328.2 K and p = 0.1 MPa:	https://www.doi.org/10.1016/j.jct.2011.05.022
Isobaric vapor-liquid equilibrium for binary system of 2-ethylthiophene + n-octane at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2014.07.010
Thermodynamic Properties of Mixtures Containing Precursors of Vitamin B5:	https://www.doi.org/10.1021/je020060f

Thermodynamic Properties of Mixtures Containing Ionic Liquids. 9. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in Trimethyl-butylammonium Bis(trifluoromethylsulfonyl) Imide Using Gas-Liquid Chromatography and Static Method:	https://www.doi.org/10.1021/je050440b
Densities and Kinematic Viscosities of a Quinary Regular Liquid System and Its Five Quaternary Subsystems at 298.15 K:	https://www.doi.org/10.1007/s10765-014-1759-1
Measurement and Correlation of Diffusion Coefficients of Aromatic Compounds at Infinite Dilution in Alkane and Cycloalkane Solvents:	https://www.doi.org/10.1021/je6005604
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.01.019
Vapor-Liquid Equilibria for Binary Mixtures: 1. Ethanol + 1-Butanol, Ethanol + Octane, 1-Butanol + Octane:	https://www.doi.org/10.1021/je0600045
Vapor Liquid Equilibria Measurements for the Nine n-Alkane/Ketone Pairs Comprising 2-, 3-, and 4-Heptanone with n-Octane, n-Nonane, and n-Decane:	https://www.doi.org/10.1021/je500731x
Activity coefficients at infinite dilution for solutes in the trioctylmethylammonium bis(trifluoromethylsulfonyl)imide ionic liquid using gas liquid chromatography:	https://www.doi.org/10.1016/j.jct.2009.08.012
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.02.012
(p, rho, T, x) and viscosity measurements of {x1 n-heptane + (1 - x1) n-octane} mixtures at high temperatures and high pressures:	https://www.doi.org/10.1016/j.jct.2006.01.012
Novel liquid-liquid equilibrium data for six ternary systems containing IL, hydrocarbon and thiophene at 25 C:	https://www.doi.org/10.1016/j.fluid.2015.12.006
liquid-liquid equilibrium of substituted perfluoro-n-octane + n-octane systems:	https://www.doi.org/10.1016/j.fluid.2008.04.011
Activity coefficients at infinite dilution measurements for organic solutes and water in the 1-hexyloxymethyl-3-methyl-imidazolium and 1,3-dihexyloxymethyl-imidazolium bis(trifluoromethylsulfonyl)imide ionic liquids - The cation influence:	https://www.doi.org/10.1016/j.fluid.2009.08.017
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-hexyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.11.025
Liquid-Liquid Equilibria of Binary, Ternary and Quaternary Systems Formed by Ethylene Glycol, Water, Toluene, p-Xylene, Hexane, and Octane:	https://www.doi.org/10.1021/acs.jced.6b01048
Coexistence curves and critical behavior of nonaqueous microemulsions of (AOT + DMA + nonane) and (AOT + DMA + octane):	https://www.doi.org/10.1016/j.jct.2005.03.010
Infinite Dilution Activity Coefficients of Solutes Dissolved in Two Trihexyl(tetradecyl)phosphonium Ionic Liquids:	https://www.doi.org/10.1021/je500050p
Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je800658v
Activity coefficients at infinite dilution, physicochemical and thermodynamic properties of organic solutes and water in the ionic liquid ethyl-dimethyl-(2-methoxyethyl)ammonium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2015.05.022
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium 2-(2-methoxyethoxy) ethylsulfate at T = (308.15, 313.15, 323.15 and 333.15) K using gas + liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.017
Experimental and Nitta-Chao model prediction of high pressure density of p-xylene with dialkyl carbonates or n-alkanes:	https://www.doi.org/10.1016/j.jct.2013.09.039
Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Excess Molar Volumes for Three and Four Component Mixtures Simulating the Binary Mixture (Cyclohexane+ Hexadecane):	https://www.doi.org/10.1021/acs.jced.5b00331
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C111659&Units=SI
Vapor-liquid equilibrium in the binary and ternary systems containing ethyl propionate, ethanol and alkane at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2018.07.013
Oil desulfurization using deep eutectic solvents as sustainable and economical extractants via liquid-liquid extraction: Experimental and PC-SAFT predictions:	https://www.doi.org/10.1016/j.fluid.2018.03.018
The study of activity coefficients at infinite dilution for organic solutes and water in 1-butyl-4-methylpyridinium dicyanamide, [B4MPy][DCA] using GLC:	https://www.doi.org/10.1016/j.jct.2013.09.007
The ionic liquid 1-ethyl-3-methylimidazolium tris(pentafluoroethyl)trifluorophosphate as alternative extractant for BTEX separation:	https://www.doi.org/10.1016/j.fluid.2015.06.046
Activity coefficients at infinite dilution measurements for organic solutes in the ionic liquid trihexyltetradecylphosphonium bis(2,4,4-trimethylpentyl)-phosphinate using g.l.c. at T = (303.15, 308.15, 313.15) K:	https://www.doi.org/10.1016/j.jct.2008.04.002
Measurement of vapor liquid equilibria (VLE) and excess enthalpies (HE) of binary systems of 1-alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide and prediction of these properties and γ∞ using modified UNIFAC (Dortmund):	https://www.doi.org/10.1016/j.fluid.2007.06.001
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in 1-Propyl-1-methylpiperidinium Bis(trifluoromethylsulfonyl)imide Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je800043a
Liquid-liquid equilibria for ternary systems acetic acid + n-butylacetate + hydrocarbons at 293.15, 303.15 and 313.15 K:	https://www.doi.org/10.1016/j.fluid.2013.07.032
Excess enthalpies of binary mixtures of 2-ethoxyethanol with four hydrocarbons at 298.15, 308.15, and 318.15K: An experimental and theoretical study:	https://www.doi.org/10.1016/j.fluid.2006.04.016
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 4-methyl-N-butyl-pyridinium bis(trifluoromethylsulfonyl)-imide:	https://www.doi.org/10.1016/j.jct.2009.06.011
The Critical Temperatures of a Number of (i) (Chloroalkane (C3,C4) + Hydrocarbon (C6-C15)) Binary Mixtures and (ii) (Aromatic Halocarbon (Chlorobenzene, Fluorobenzene, 1,2-Dichlorobenzene, or 1,3-Dichlorobenzene) + Alkane (C8)) Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.7b00191
Solution thermodynamics near the liquid-liquid critical point. I. First-order corrections:	https://www.doi.org/10.1016/j.fluid.2007.05.017
Thermodynamics of mixtures containing amines VI. Liquid liquid equilibria for mixtures of o-toluidine + selected alkanes:	https://www.doi.org/10.1016/j.fluid.2007.01.022
Excess Molar Volumes and Viscosities of Binary Systems of Butylcyclohexane with n-Alkanes (C7 to C14) at T = 293.15 K to 313.15 K:	https://www.doi.org/10.1021/je400835u
Vapor Liquid Equilibria Measurements for the Five Linear C6 Esters with n-Octane:	https://www.doi.org/10.1021/acs.jced.5b01058
Miscibility of Trihexyl(tetradecyl)phosphonium Chloride with Alkanes:	https://www.doi.org/10.1021/je9010079
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-octyl-3-methylimidazolium hexafluorophosphate using gas liquid chromatography at T = (313.15, 323.15, and 333.15) K:	https://www.doi.org/10.1016/j.jct.2009.12.004
KDB Pure (Korean Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=8
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in 1-Butyl-3-methylimidazolium Dibutylphosphate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je201310d
Infinite dilution activity coefficients, specific retention volumes and solvation thermodynamics of hydrocarbons in C78H158 branched alkane solvent:	https://www.doi.org/10.1016/j.fluid.2006.07.015

Solubility of Anthracene in Binary Diisopropyl Ether + Alkane Solvent Mixtures at 298.15 K.	https://www.doi.org/10.1021/je060380p
Measurements of Activity Coefficients at Infinite Dilution for Organic Solutes and Water in the Ionic Liquid 1-Butyl-1-methylpiperidinium Thiocyanate.	https://www.doi.org/10.1021/je101008y
Physicochemical properties and activity coefficients at infinite dilution for organic solutes and water in a novel [illegible]	https://www.doi.org/10.1016/j.jct.2012.09.033
Thermodynamic behaviour of the binary systems dimethyl carbonate + [illegible]	https://www.doi.org/10.1016/j.tca.2004.11.022
Experimental Study of Thermodynamic Properties of Mixtures Containing Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate Using Gas-Liquid Chromatography and Transpiration Method	https://www.doi.org/10.1021/je0602723
Enhancing liquid-phase olefin-paraffin separations using novel silver-based [illegible]	https://www.doi.org/10.1021/je5004669
Viscosity and Density of Five Hydrocarbon Liquids at Pressures up to 200 MPa and Temperatures up to 473 K.	https://www.doi.org/10.1021/je800417q
Thermodynamic Properties of Binary Mixtures of p-Xylene with Cyclohexane, Heptane, Octane, and N-Methyl-2-pyrrolidone at Several Temperatures.	https://www.doi.org/10.1021/je049776w
Activity coefficients at infinite dilution of organic solutes in the ionic liquid [illegible]	https://www.doi.org/10.1016/j.jct.2009.07.010
Solubility of 2-Ethylanthraquinone in [illegible]	https://www.doi.org/10.1021/acs.jced.6b00334
Excess Enthalpies of Binary Mixtures [illegible]	https://www.doi.org/10.1021/je030215z
[illegible] binary mixtures of dicationic ionic liquids for determination of activity coefficients at infinite dilution [illegible]	https://www.doi.org/10.1016/j.fluid.2013.05.038
Separation of aliphatic from aromatic hydrocarbons and sulphur compounds from fuel based on measurements of activity coefficients at infinite dilution [illegible]	https://www.doi.org/10.1016/j.jct.2016.07.017
Densities and Viscosities of Corn Oil + n-Alkanes Blends from (288.15 to [illegible]	https://www.doi.org/10.1021/acs.jced.7b00121
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl [illegible]	https://www.doi.org/10.1021/je6005696
Thermodynamics of solvation in N,N-diethyl-N-methyl-N-(2-methoxy-ethyl)ammonium [illegible]	https://www.doi.org/10.1016/j.jct.2014.06.006
[illegible] tetrafluoroborate Using Gas-Liquid Chromatography	https://www.doi.org/10.1021/je0302295
Comparative Tensimetry: 1-Propanol and 1-Butanol in Heptane and in Octane.	https://www.doi.org/10.1021/je200637v
(Liquid + liquid) equilibrium of binary mixtures formed by N,N-dimethyl [illegible]	https://www.doi.org/10.1016/j.fluid.2007.02.026
Densities and derived thermodynamic properties for the (n-heptane + n-octane), (n-heptane + ethanol) and (n-octane + ethanol) systems at high pressures	https://www.doi.org/10.1016/j.jct.2016.04.014
Activity coefficients at infinite dilution of organic solvents and water in [illegible]	https://www.doi.org/10.1016/j.fluid.2016.02.004
Measurements of activity coefficients [illegible]	https://www.doi.org/10.1016/j.jct.2010.10.026
Measurements of activity coefficients at infinite dilution of organic solutes [illegible]	https://www.doi.org/10.1016/j.jct.2012.03.005
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=8
Comparison of Extractability of Oxalic Acid from Dilute Aqueous Solutions Using Dioctylamine and [illegible]	https://www.doi.org/10.1021/acs.jced.8b01155
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic [illegible]	https://www.doi.org/10.1016/j.jct.2018.01.003
Viscosity of the binary systems 2-methyl-2-propanol with n-alkanes at T = (303.15, 308.15, 313.15, 318.15 and 323.15) K: Prediction and correlation - New UNIFAC-VISCO interaction [illegible]	https://www.doi.org/10.1016/j.fluid.2010.10.009
Dynamic Viscosities of 2-Butanol with Alkanes (Octane, Decane, and [illegible]	https://www.doi.org/10.1021/je034208m
Preferential solvation and mixing [illegible]	https://www.doi.org/10.1016/j.fluid.2016.08.035
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)- [illegible]	https://www.doi.org/10.1016/j.jct.2013.01.007
Thermodynamic behavior of thiophene with octane [illegible]	https://www.doi.org/10.1016/j.jct.2015.12.031
Solubility of Benzophenone in Binary Alkane + Carbon Tetrachloride Solvent [illegible]	https://www.doi.org/10.1021/je0340497
[illegible]	https://www.doi.org/10.1021/je900838a
(Liquid + Liquid) Equilibria for the Ternary Systems of FC3283 + HFE7300 + Hexane, FC3283 + HFE7300 + Octane, and FC72 + HFE7100 + Acetonitrile [illegible]	https://www.doi.org/10.1021/je500107y
Liquid–liquid extraction of aromatics from their mixtures with alkanes using [illegible]	https://www.doi.org/10.1016/j.jct.2012.05.006
Determination of Activity Coefficients [illegible]	https://www.doi.org/10.1021/je030187k
[illegible]	https://www.doi.org/10.1016/j.fluid.2011.05.002
[illegible]	https://www.doi.org/10.1021/acs.jced.7b00869
The density, the refractive index and the adjustment of the excess thermodynamic properties by means of the multiple linear regression method [illegible]	https://www.doi.org/10.1016/j.tca.2015.08.023
Activity Coefficients at Infinite Dilution by GLC in Alkanediamines as [illegible]	https://www.doi.org/10.1021/je200628n
Measurement and prediction of activity coefficients at infinite dilution [illegible]	https://www.doi.org/10.1016/j.fluid.2008.11.013
Vapor-Liquid Equilibria (VLE) and [illegible]	https://www.doi.org/10.1021/acs.jced.6b00116
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Butyl-3-methylimidazolium Nitrate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je200050q

Phase equilibria of 1-hexyl-3-methylimidazolium acetate with water and oil:	https://www.doi.org/10.1016/j.fluid.2018.11.010
Measurement of infinite dilution activity coefficients of n-alkanes in 4-methyl-n-butylpyridinium tetrafluoroborate using gas liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2006.10.016
Liquid-Liquid(-Liquid) Equilibria in Ternary Systems of Water + Cyclohexylamine + Aromatic Hydrocarbon (Toluene or Propylbenzene) or Aliphatic Hydrocarbon (Heptane or Octane):	https://www.doi.org/10.1021/je050520f
Interactions of volatile organic compounds with the ionic liquids 1-butyl-1-methylpyrrolidinium tetracyanoborate and 1-butyl-1-methylpyrrolidinium bis(oxalato)borate:	https://www.doi.org/10.1016/j.jct.2012.09.017
Evaluation of ionic liquids as solvent for aromatic extraction: Experimental, correlation and COSMO-RS predictions:	https://www.doi.org/10.1016/j.jct.2013.07.011
Measurement of Volumetric Properties of the Ternary Systems Hexane + Heptane + Octane and Octane + Nonane + Decane Using a Flow Apparatus:	https://www.doi.org/10.1021/je501134s
Study of molecular interactions in binary liquid mixtures of 1-octanol with n-hexane, n-octane, and n-decane using volumetric, viscometric, and acoustic properties:	https://www.doi.org/10.1016/j.jct.2008.02.005
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate [EMIM][TCB] using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2010.02.006
Thermodynamics and limiting activity coefficients measurements for organic solutes and water in the ionic liquid 1-dodecyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide:	https://www.doi.org/10.1016/j.jct.2016.08.008
Separation of water/butan-1-ol based on activity coefficients at infinite dilution in 1,3-didecyl-2-methylimidazolium dicyanamide ionic liquid:	https://www.doi.org/10.1016/j.jct.2017.10.003
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Butyl-3-methyl-pyridinium Trifluoromethanesulfonate:	https://www.doi.org/10.1021/je1000582
Measurement and Modeling of Vapor Liquid Equilibria for the Octane + Ethanol System:	https://www.doi.org/10.1021/je400238j
Liquid phase behaviour of 1-hexyloxymethyl-3-methyl-imidazolium-based ionic liquids with hydrocarbons: The influence of anion:	https://www.doi.org/10.1016/j.jct.2005.02.012
Experimental and predicted phase equilibria and excess properties for systems with ionic liquids:	https://www.doi.org/10.1016/j.fluid.2008.03.011
Activity Coefficients at Infinite Dilution for Organic Compounds Dissolved in 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Having Six-, Eight-, and Ten-Carbon Alkyl Chains:	https://www.doi.org/10.1021/je300692s
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in 1-allyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ionic liquid:	https://www.doi.org/10.1016/j.jct.2016.06.028
Phase Equilibria of (1-Hexyl-3-methylimidazolium Thiocyanate + Water, Alcohol, or Hydrocarbon) Binary Systems:	https://www.doi.org/10.1021/je900460m
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Hexyl-3-methylimidazolium Hexafluorophosphate from Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je0256481
Activity coefficients at infinite dilution measurements for organic solutes in the ionic liquid 1-hexyl-3-methyl-imidazolium bis(trifluoromethylsulfonyl)-imide using g.l.c. at T = (298.15, 313.15, and 333.15) K:	https://www.doi.org/10.1016/j.jct.2005.03.014
Densities and Kinematic Viscosities of One Quinary Regular Liquid System at Temperatures from (288.15 to 298.15) K:	https://www.doi.org/10.1007/s10765-012-1159-3
Viscosities and Interfacial Properties of 1-Methyl-3-butylimidazolium Hexafluorophosphate and 1-Isobutenyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids:	https://www.doi.org/10.1021/je0604277
The measurements of coexistence curves and light scattering for [illegible] in the critical region:	https://www.doi.org/10.1016/j.jct.2008.07.008
Activity Coefficients at Infinite Dilution of Organic Compounds in Trihexyl(tetradecyl)phosphonium Bis(trifluoromethylsulfonyl)imide Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je800754w
Deduction of Physicochemical Properties from Solubilities: 2,4,5-Trimethoxybenzoic Acid, Biotin, and Caprolactam as Examples:	https://www.doi.org/10.1021/je501140p
Extraction of toluene from aliphatic compounds using an ionic liquid as solvent: Influence of the alkane on the (liquid + liquid) equilibrium:	https://www.doi.org/10.1016/j.jct.2010.11.013
Density, Viscosity, and Speed of Sound of (Methyl Benzoate + Cyclohexane), (Methyl Benzoate + n-Hexane), (Methyl Benzoate + Heptane), and (Methyl Benzoate + Octane) at Temperatures of (293.15, 303.15, and 313.15) K:	https://www.doi.org/10.1021/je9006597
Apparent and Partial Molar Volumes at Infinite Dilution and Solid-Liquid Equilibria of Dibenzothiophene + Alkane Systems (298.15, 308.15, and 313.15) K:	https://www.doi.org/10.1021/je200327s
Experimental determination of critical data of multi-component mixtures containing potential gasoline additives 2-butanol by a flow-type apparatus:	https://www.doi.org/10.1016/j.jct.2016.05.010
Thermodynamic Studies of Solubility for Naphthalene in 12 Solvents from 273 to 330 K:	https://www.doi.org/10.1021/acs.jced.9b00064
Heat capacity and Joule-Thomson coefficient of selected n-alkanes at 0.1 and 10 MPa in broad temperature ranges:	https://www.doi.org/10.1016/j.jct.2017.03.034
Infinite dilution activity coefficients of volatile organic compounds in two ionic liquids composed of the tris(pentafluoroethyl)trifluorophosphate ([FAP]) anion and a functionalized cation:	https://www.doi.org/10.1016/j.jct.2013.05.035
Separation of Benzene from Linear Alkanes (C6-C9) Using 1-Ethyl-3-Methylimidazolium Ethylsulfate at T = 298.15 K:	https://www.doi.org/10.1021/je1001544
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Hexyl-3-methylimidazolium Thiocyanate:	https://www.doi.org/10.1021/je900890u
Extraction of Benzene from Aliphatic Compounds Using Commercial Ionic Liquids as Solvents: Study of the Liquid-Liquid Equilibrium at T = 298.15 K:	https://www.doi.org/10.1021/je200334e
Thermodynamics and activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.06.007
Excess properties of the binary mixtures of methylcyclohexane + alkanes (C6 to C12) at T = 298.15 K to T = 308.15 K:	https://www.doi.org/10.1016/j.jct.2005.03.024
Experimental study on liquid-liquid equilibria of ionic liquids + alkane + ethyl benzene or p-xylene at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2015.08.031
Liquid-Liquid Equilibria for the Binary Systems of Sulfolane with Alkanes:	https://www.doi.org/10.1021/je7001607
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on gamma infinity data measurements:	https://www.doi.org/10.1016/j.jct.2017.11.017
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je700282w
Densities and Viscosities of MTBE + Heptane or Octane at p)0.1 MPa from (273.15 to 363.15) K:	https://www.doi.org/10.1021/je600554h
Measurements of activity coefficients at infinite dilution of aromatic and aliphatic hydrocarbons, alcohols, and water in the new ionic liquid [EMIM][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2008.01.004
Effect of bromine substitution on the solubility of gases in hydrocarbons and fluorocarbons:	https://www.doi.org/10.1016/j.fluid.2006.11.013

Thermodynamics of Mixtures Containing a Strongly Polar Compound. 8. Liquid-Liquid Equilibria for N,N-Dialkylamide + Selected N-Alkanes:	https://www.doi.org/10.1021/je050428j
Liquid-Liquid Equilibria of Methyl Acetate + Methanol + Octane or Nonane:	https://www.doi.org/10.1021/je0342365
Viscosity measurements of hydrocarbon fuel at temperatures from (303.2 to 513.2) K and pressures up to 5.1 MPa using a two-capillary viscometer:	https://www.doi.org/10.1016/j.tca.2015.08.005
Densities, Viscosities, and Refractive Indices of Binary Mixtures of 1,2,3,4-Tetrahydronaphthalene with Some n-Alkanes at T = (293.15 to 313.15) K:	https://www.doi.org/10.1021/je300899n
Measurement of activity coefficients at infinite dilution in 1-hexadecyl-3-methylimidazolium tetrafluoroborate ionic liquid:	https://www.doi.org/10.1016/j.jct.2007.01.004
Vapor-liquid equilibria in ternary systems of associating components (water, aniline, cyclohexylamine) and hydrocarbons (octane or toluene):	https://www.doi.org/10.1016/j.fluid.2007.07.023
Liquid Extraction of Toluene from Heptane, Octane, or Nonane Using Mixed Ionic Solvents of 1-Ethyl-3-methylimidazolium Methylsulfate and 1-Hexyl-3-methylimidazolium Hexafluorophosphate:	https://www.doi.org/10.1021/acs.jced.8b00669
Activity coefficients at infinite dilution of organic solutes in methylphosphonate based ionic liquids using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2015.02.023
Experimental determination and calculation of the critical curves for the binary systems of CO2 containing ketone, alkane, ester and alcohol, respectively:	https://www.doi.org/10.1016/j.fluid.2011.05.005
Ternary Liquid-Liquid Equilibria for Aniline + Cyclohexylamine + Water, Aniline + Cyclohexylamine + Octane, Aniline + Water + Toluene, and Aniline + Water + Octane:	https://www.doi.org/10.1021/je0505366
Activity Coefficients at Infinite Dilution of Various Solutes in Tetrapropylammonium Bromide + 1,6-Hexanediol Deep Eutectic Solvent:	https://www.doi.org/10.1021/acs.jced.8b00600
Phase Equilibria of (1-Ethyl-3-methylimidazolium Ethylsulfate + Hydrocarbon, + Ketone, and + Ether) Binary Systems:	https://www.doi.org/10.1021/je700591h
High selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid:	https://www.doi.org/10.1016/j.fluid.2017.06.001
Liquid-Liquid Equilibria of Binary and Ternary Systems Methanol/Water + n-Hexane, + n-Octane, + n-Dodecane, and + n-Hexadecane in the Temperature Range between (T) = 283.15 K and T = 333.15 K:	https://www.doi.org/10.1021/acs.jced.6b00022
Excess molar volumes of the ternary system {methylcyclohexane (1) + cyclohexane (2) + n-alkanes (3)} at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2005.02.006
Liquid-liquid extraction of p-xylene from their mixtures with alkanes using 1-butyl-1-methylmorpholinium tricyanomethanide and 1-butyl-3-methylimidazolium tricyanomethanide ionic liquids:	https://www.doi.org/10.1016/j.fluid.2015.12.032
Ternary Liquid-Liquid Equilibria for Systems of (Sulfolane + Toluene or Chlorobenzene + Octane):	https://www.doi.org/10.1021/je0601210
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-butyl-3-methylimidazolium hexafluoroantimonate using gas-liquid chromatography at T = (313.15, 323.15, and 333.15) K:	https://www.doi.org/10.1016/j.jct.2010.12.019
Liquid-liquid equilibria for (2-hydroxy benzaldehyde + n-alkane) mixtures. Intermolecular and proximity effects in systems containing hydroxyl and aldehyde groups:	https://www.doi.org/10.1016/j.jct.2019.04.002
Activity Coefficients at Infinite Dilution of Organic Compounds in Four New Imidazolium-Based Ionic Liquids:	https://www.doi.org/10.1021/je200195q
Density, Surface Tension, and Refractive Index of Octane + 1-Alkanol Mixtures at T) 298.15 K:	https://www.doi.org/10.1021/je034053i
Activity coefficients at infinite dilution for organic solutes and water in 1-ethyl-1-methylpyrrolidinium lactate:	https://www.doi.org/10.1016/j.jct.2015.05.014
Simultaneous Measurement of Dynamic Viscosity and Density of n-Alkanes at High Pressures:	https://www.doi.org/10.1021/acs.jced.7b00650
Diffusion Coefficients of Carbon Dioxide in Eight Hydrocarbon Liquids at Temperatures between (298.15 and 423.15) K at Pressures up to 69 MPa:	https://www.doi.org/10.1021/acs.jced.6b00691
Comparative study of the LLE of the quaternary and ternary systems involving benzene, n-octane, n-decane and the ionic liquid [EMpy][NTf2]:	https://www.doi.org/10.1016/j.jct.2016.02.027
Measurements of activity coefficients at infinite dilution of aliphatic and aromatic hydrocarbons, alcohols, thiophene, tetrahydrofuran, MTBE, and water in ionic liquid [BMIM][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2008.12.018
Fully automatized apparatus for determining speed of sound for liquids in the temperature and pressure interval (283.15-343.15) K and (0.1-95) MPa:	https://www.doi.org/10.1016/j.jct.2016.09.022
Liquid-liquid equilibrium (LLE) data for ternary mixtures of {aliphatic + p-xylene + [EMpy][ESO4]} at T = 313.15 K:	https://www.doi.org/10.1016/j.fluid.2012.07.005
Measurements of Density and Heat Capacity for Binary Mixtures {x Benzonitrile + (1 - x) (Octane or Nonane)}:	https://www.doi.org/10.1021/je100145f
Liquid-Liquid Equilibria Data for Ethylbenzene or p-Xylene with Alkane and 1-Butylpyridinium Nitrate Ionic Liquid at 298.15 K:	https://www.doi.org/10.1021/acs.jced.6b00881
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.05.017
Measurements of different thermodynamic properties of systems containing ionic liquids and correlation of these properties using modified UNIFAC (Dortmund):	https://www.doi.org/10.1016/j.fluid.2010.02.010
p-(Vm)-T measurements of (octane + benzene) at temperatures from (298.15 to 328.15) K and at pressures up to 40 MPa:	https://www.doi.org/10.1016/j.jct.2007.11.009
Liquid-Liquid Equilibrium Data for Ternary Systems Containing Alkanes (n-Pentane, n-Hexane, n-Heptane, and n-Octane) + Alcohol (Methanol and Ethanol) + Protic Ionic Liquid (2-HEAF):	https://www.doi.org/10.1021/acs.jced.9b00375
Thermodynamics of Mixtures Containing Aromatic Alcohols. 1. Liquid-Liquid Equilibria for (Phenylmethanol + Alkane) Systems:	https://www.doi.org/10.1021/je300146k
Isothermal and isobaric (vapour + liquid) equilibria of (alpha-pinene + n-butanol + n-octane):	https://www.doi.org/10.1016/j.jct.2004.08.012
Measurements of the Thermal Conductivity of n-Octane, Isooctane, 1-Octene, and 1-Octanol in the Temperature Range from 253 to 393 K at Pressures up to 30 MPa:	https://www.doi.org/10.1021/acs.jced.9b00628
Separation of binary mixtures based on limiting activity coefficients data using specific ammonium-based ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.fluid.2017.12.029
Extraction of Citric Acid and Malic Acid from Their Aqueous Solutions Using a Phosphorus-Bonded Extractant, Tri-n-octylphosphineoxide, and a Secondary Amine, Dioctylamine:	https://www.doi.org/10.1021/acs.jced.7b00562
Study of critical and maximum temperatures of coexistence of liquid and gas phases in hydrocarbons binary mixtures of aromatic hydrocarbons with alkanes and cycloalkanes:	https://www.doi.org/10.1016/j.fluid.2014.06.020
Vapor-liquid equilibrium data for the nitrogen + n-octane system from (344.5 to 543.4) K and at pressures up to 50 MPa:	https://www.doi.org/10.1016/j.fluid.2009.04.015
Determination and correlation of liquid liquid equilibria for four binary N,N-dimethylformamide + hydrocarbon systems:	https://www.doi.org/10.1016/j.fluid.2007.03.010
Measurement and Prediction of Excess Properties of Binary Mixtures Methyl Decanoate + an Even-Numbered n-Alkane (C6-C16) at 298.15 K:	https://www.doi.org/10.1021/acs.jced.9b00185

Measurements of activity coefficients at infinite dilution of organic solutes and water in ...	https://www.doi.org/10.1016/j.jct.2010.05.017
Extraction of toluene from alkane using [Bmim][NO3] or [Omim][NO3] ionic liquids ...	https://www.doi.org/10.1016/j.fluid.2013.11.020
Viscosities of Dimethyl Carbonate or Diethyl Carbonate with Alkanes at Four Temperatures. New UNIFAC-VISCO Parameters	https://www.doi.org/10.1021/je020131a
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid ...	https://www.doi.org/10.1016/j.jct.2015.02.024
Analysis of Surface Tension, Density, and Speed of Sound for the Ternary ...	https://www.doi.org/10.1021/je800594e
Activity coefficients at infinite dilution of organic solutes in the ionic liquid trihexyltetradecylphosphonium ...	https://www.doi.org/10.1016/j.jct.2013.05.030
Viscosity ... at 243 ... 353 K ...	https://www.doi.org/10.1021/acs.jced.7b00396
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... parameters of ionic liquid ...	https://www.doi.org/10.1016/j.jct.2018.09.023
Abraham model linear free energy relationships for describing the partitioning and solubility behavior of ...	https://www.doi.org/10.1016/j.fluid.2016.10.009
Determination of Activity Coefficients at Infinite Dilution of Organic Solutes in ...	https://www.doi.org/10.1021/acs.jced.8b00080
Excess molar enthalpies of the ternary mixtures: (tetrahydrofuran or ...) at 298.15 K	https://www.doi.org/10.1016/j.jct.2005.07.017
... 1-(2-Hydroxyethyl)-3-methylimidazolium ...	https://www.doi.org/10.1021/je034182z
Determination of Activity Coefficients at Infinite Dilution of Solutes in ...	https://www.doi.org/10.1021/acs.jced.8b00635
Phase Equilibria Study of the Binary Systems (N-Butyl-4-methylpyridinium ...)	https://www.doi.org/10.1021/je8006309
Screening of environmental friendly ... solvent for the different types of separations problem: ...	https://www.doi.org/10.1016/j.jct.2015.08.017
Interactions of Volatile Organic Compounds with the Ionic Liquid ...	https://www.doi.org/10.1021/je200822w
Thermodynamic Parameters of a New Synthesized Tricationic Ionic Liquid Stationary Phase by Inverse Gas Chromatography	https://www.doi.org/10.1021/acs.jced.8b00601
... Ternary Systems of 4-Methyl-1,3-dioxolan-2-one + 1,4-Dimethylbenzene + Octane, ...	https://www.doi.org/10.1021/je401056v
Experimental and theoretical study on infinite dilution activity coefficients of ...	https://www.doi.org/10.1016/j.jct.2013.01.005
Liquid–liquid equilibrium data for binary systems of ... perfluoroalkane (C6 and C8) + n-alkane ...	https://www.doi.org/10.1016/j.fluid.2010.05.018
... Vibrating Edge Supported Plate, Fabricated by the Methods of Micro Electro Mechanical System for the Simultaneous Measurement of Density and Viscosity: Results for (n-heptane, ...)	https://www.doi.org/10.1021/je0503296
Density, viscosity and excess molar volume of binary mixtures of ...	https://www.doi.org/10.1016/j.jct.2013.09.017
Liquid–Liquid Equilibria of the Ternary Systems of Alkane + Aromatic + ...	https://www.doi.org/10.1021/je100716c
...	https://www.doi.org/10.1016/j.jct.2005.01.016
Measurement and prediction of pyrene solubility in pure, binary, ternary and quaternary solvent systems	https://www.doi.org/10.1016/j.fluid.2007.11.001
Excess Enthalpies and Thermal Conductivity Coefficients for Binary Mixtures of Carbon Tetrachloride and Four Alkanes (C5 to C8) ...	https://www.doi.org/10.1021/je7007395
Activity coefficients at infinite dilution of organic solutes in diethylene glycol ...	https://www.doi.org/10.1016/j.jct.2013.05.011
Dynamic viscosities of butanol with alkanes (C8, C10, and C12) at several temperatures	https://www.doi.org/10.1016/j.jct.2003.12.005
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in ...	https://www.doi.org/10.1021/je060142u
Surface tension and density of mixtures of m-xylene + n-alkane at ...	https://www.doi.org/10.1016/j.jct.2016.11.018
Liquid–Liquid Extraction of Toluene from ...	https://www.doi.org/10.1021/je5001439
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Design of a Gamma Densitometer for hydrocarbon Fuel at High Temperature and Supercritical Pressure	https://www.doi.org/10.1021/je500314x
Interactions of volatile organic compounds with the ionic liquid 1-ethyl-3-methylimidazolium ...	https://www.doi.org/10.1016/j.jct.2011.09.028
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Activity coefficients at infinite dilution of organic solutes in N-formylmorpholine and N-methylpyrrolidone from gas-liquid chromatography	https://www.doi.org/10.1016/j.jct.2013.02.006
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Volumetric and Transport Properties of Binary Mixtures of n-Octane + Ethanol, ... 1-Pentanol from (293.15 to 323.15) K at Atmospheric Pressure	https://www.doi.org/10.1021/je4004806
Liquid–liquid extraction of toluene from alkane with pyridinium-based ionic ...	https://www.doi.org/10.1016/j.jct.2016.07.030
Ammonium ionic liquids in extraction of ... from water phase using activity coefficients at infinite dilution	https://www.doi.org/10.1016/j.fluid.2018.09.024
Phase Separation in Binary Mixtures Containing Linear Perfluoroalkanes: ...	https://www.doi.org/10.1021/je049620w

(Solid + liquid) and (liquid + liquid) phase equilibria measurements and correlation of the binary systems	https://www.doi.org/10.1016/j.fluid.2009.01.012
Measurement and Correlation of High Pressure Phase Equilibria for CO2 + Isosorbide + Alcohol or Hydrocarbon):	https://www.doi.org/10.1021/acs.jced.7b00517
Vapor-Liquid Equilibria for the Binary Mixture alpha-Pinene + Octane:	https://www.doi.org/10.1021/je025631i
Study of the suitability of two ammonium-based ionic liquids for the extraction of benzene from mixtures with aliphatic hydrocarbons.:	https://www.doi.org/10.1016/j.fluid.2016.02.006
Determination and correlation of (liquid + liquid) equilibria of ternary and quaternary systems with octane +	https://www.doi.org/10.1016/j.jct.2015.11.016
Study of interaction between organic compounds and nanoparticles by thermodynamics of mixtures	https://www.doi.org/10.1016/j.fluid.2014.11.020
containing amines. XI. Liquid + liquid equilibria and molar excess enthalpies at 298.15 K for N-methylaniline +	https://www.doi.org/10.1016/j.jct.2012.07.006
Liquid Viscosity and Surface Tension of Hexane, Octane, Decane, and Hexadecane up to 573 K by Surface	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Light Scattering	https://www.doi.org/10.1021/acs.jced.9b00525
Activity coefficients at infinite dilution measurements for organic solutes in the ionic liquid	https://www.doi.org/10.1016/j.jct.2005.01.015
Comparison Study on Temperature Dependence of the Interfacial Tension of Toluene or Octane +	https://www.doi.org/10.1021/je500450w
1-Butyl-3-methylimidazolium 2-(2-methoxyethoxy) ethyl sulfate using	https://www.doi.org/10.1016/j.jct.2018.05.017
Activity coefficients at infinite dilution measurements for octane +	https://www.doi.org/10.1016/j.jct.2005.09.003
Activity coefficients at infinite dilution of alkanes (hexane, heptane, and octane) in the ionic liquid	https://www.doi.org/10.1021/je1007923
Ternary Systems of Aliphatic Hydrocarbons (Heptane or Octane) +	https://www.doi.org/10.1016/j.jct.2008.09.004
Activity coefficients at infinite dilution of hydrocarbons, alkylbenzenes, and alcohols in the pharmaceutical	https://www.doi.org/10.1016/j.fluid.2015.05.031
Experimental Measurements and Modeling of Carbon Dioxide Gas Solubility and	https://www.doi.org/10.1016/j.jct.2016.11.017
Activity Coefficient Modeling for Propan-1-ol + n-Octane or n-Nonane or n-Decane from 313.15 K to 363.15 K	https://www.doi.org/10.1021/acs.jced.8b00554
Isobaric Vapor-Liquid Equilibrium of Binary Systems of Hexane or Octane	https://www.doi.org/10.1021/je3002283
Activity coefficients at infinite dilution of alkanes in 2,2,4-trimethylpentane and	https://www.doi.org/10.1016/j.jct.2014.12.023
Activity Coefficients at Infinite Dilution for Alkanes, Esters, and Ethers in Ionic Liquid Based on	https://www.doi.org/10.1021/acs.jced.5b00813
Physicochemical Behavior of Binaries and Ternaries of Octane or	https://www.doi.org/10.1016/j.jct.2011.04.005
Octane with Methyl Esters (Ethanoate, Butanoate, Pentanoate):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=8
Determination of activity coefficients at infinite dilution of organic solutes in	https://www.doi.org/10.1016/j.fluid.2005.04.021
Activity coefficients at infinite dilution measurements for organic solutes and	https://www.doi.org/10.1016/j.jct.2008.12.005
Physical properties of anisole +	https://www.doi.org/10.1016/j.jct.2004.09.021
Liquid-Liquid Equilibria for Propylene Glycol Hydrocarbon Binary Systems:	https://www.doi.org/10.1021/je3009039
Thermodynamics and activity coefficients at infinite dilution for	https://www.doi.org/10.1016/j.jct.2018.08.028
Estimated Solubility Method: organic compounds in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Activity Coefficients at Infinite Dilution of Cylcohexylamine + Octane, Toluene,	https://www.doi.org/10.1021/je050316s
Activity coefficients at infinite dilution and physicochemical properties for	https://www.doi.org/10.1016/j.jct.2011.04.018
Molar volumes in binary mixtures of octane + 1-alkanol at 298.15 K	https://www.doi.org/10.1016/j.tca.2015.11.024
Excess Molar Enthalpies and Heat Capacities of {2-Methylpiperidine	https://www.doi.org/10.1021/je5008444
Activity coefficients at infinite dilution of organic solutes in the ionic liquid	https://www.doi.org/10.1016/j.jct.2010.12.005
Partition Coefficients of Organic Compounds in New Imidazolium and	https://www.doi.org/10.1021/je9003178
Measurement of the Speed of Sound in Liquid and Supercritical n-Octane	https://www.doi.org/10.1021/acs.jced.7b00712
Solubility of Gases and Liquids at Infinite Dilution in the Gaseous and Liquid	https://www.doi.org/10.1021/je700297c
Solutes and Their Thermodynamics of Solubilization in the Novel	https://www.doi.org/10.1016/j.fluid.2005.05.003
Experimental determination and calculation of thermodynamic properties of CO2 + octane + squalane containing binary systems:	https://www.doi.org/10.1016/j.fluid.2006.05.014
Perfluoroalkanes and perfluoroalkylalkane surfactants in solution: Partial molar volumes in	https://www.doi.org/10.1016/j.fluid.2008.02.008
Solubility of dibenzothiophene in nine organic solvents: Experimental measurement and thermodynamic modelling:	https://www.doi.org/10.1016/j.fluid.2011.02.020
	https://www.doi.org/10.1016/j.jct.2018.09.017

(Solid + liquid) and (liquid + liquid) phase equilibria study and correlation of the binary systems {1-butyl-3-methylpyridinium tosylate + water or + alcohols}	https://www.doi.org/10.1016/j.fluid.2010.01.025
Twenty Years of Experimental Determinations of Thermophysical Properties at the Thermodynamic Laboratory, ESIQIE-IPN, Mexico	https://www.doi.org/10.1021/acs.jced.8b01027
Liquid-Liquid Equilibrium of Ternary Systems Sulfide + Octane + Solvents at Different Temperatures	https://www.doi.org/10.1021/je800340v
Activity Coefficients at Infinite Dilution of Alkanes, Alkenes, and Alkyl Benzenes in Glycerol Using Gas-Liquid Chromatography	https://www.doi.org/10.1021/je900686n
Computer Aided Solvent screening for the Separation of binary mixtures based on gamma infinity data using McGowan Method	https://www.doi.org/10.1021/je2009889
Separation of 2-methoxynaphthene and 2-acetyl-6-methoxynaphthalene	https://www.doi.org/10.1016/j.jct.2018.09.003
[OMIM][BF4] ionic liquid and modelling of thermodynamic functions:	http://link.springer.com/article/10.1007/BF02311772
Activity coefficients at infinite dilution of organic solutes in the ionic liquid, methyltrioctylammonium thiosalicylate, [N1888][TS] by gas liquid chromatography at T = (303.15, 313.15, 323.15) K	https://www.doi.org/10.1016/j.jct.2010.12.020
Solubilities of n-Alkanes (C6 to C8) in Water from 30 °C to 130 °C	https://www.doi.org/10.1021/je025659u
Activity coefficients at infinite dilution of alkanes, alkenes, alkyl benzenes in dimethylphosphate based ionic liquids using gas liquid chromatography	https://www.doi.org/10.1016/j.jct.2015.07.046
Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K	https://www.doi.org/10.1016/j.jct.2016.02.006
High pressure densities and derived thermodynamic properties for the n-heptane + n-octane + ... ternary system	https://www.doi.org/10.1016/j.jct.2016.11.005
Determination of Activity Coefficients at Infinite Dilution of 35 Solutes in the Ionic Liquid, 1-Butyl-3-methylimidazolium Tosylate, Using Gas−Liquid Chromatography	https://www.doi.org/10.1021/je100410k
A Study on Properties Derived from Densities and Viscosities for the (Methyl Heptanoate or Ethyl Heptanoate) + n-Octane + 1-Hexanol and their Binary Subsystems	https://www.doi.org/10.1021/je034017j
Study of densities, viscosities, and speeds of sound of binary liquid mixtures	https://www.doi.org/10.1016/j.jct.2007.05.016
Excess molar volumes of (octane + benzene, toluene, or xylenes) at T = (298.15, 303.15, and 308.15) K	https://www.doi.org/10.1016/j.jct.2007.10.010
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid	https://www.doi.org/10.1016/j.jct.2012.03.015
Thermodynamics of mixtures containing aromatic nitriles:	https://www.doi.org/10.1016/j.jct.2017.09.027
Volumetric properties of dialkyl carbonate + n-alkane mixtures at high pressures: Experimental measurement and Prediction	https://www.doi.org/10.1016/j.jct.2012.11.011
1-(2-methoxyethyl)-1-methylpyrrolidinium ... Activity coefficients at infinite dilution of hydrocarbons in glycols	https://www.doi.org/10.1016/j.jct.2016.10.013
Thermodynamics and activity coefficients at infinite dilution: Experimental data and thermodynamic modeling with the GCA-EoS	https://www.doi.org/10.1016/j.fluid.2018.06.013
Activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-butyl-1-methylpyrrolidinium dicyanamide	https://www.doi.org/10.1021/je3011979
Experimental Data of the Binaries ... Activity coefficients at infinite dilution measurements for organic solutes	https://www.doi.org/10.1016/j.fluid.2008.10.008
Determination of the Critical Properties of C6-C10 n-Alkanes and Their Binary Systems Using a Flow Apparatus	https://www.doi.org/10.1021/je500756d
Mixtures of 2,2,4-Trimethylpentane + n-Heptane ... between 298.15 and 343.15 K	https://www.doi.org/10.1021/acs.jced.5b00009
Experimental Solubility Data for Binary Mixtures of Ethyl Acetate + Benzene, and Ethylbenzene	https://www.doi.org/10.1021/acs.jced.7b00613
Thermodynamic Properties of Mixtures Containing Ionic Liquids. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes	https://www.doi.org/10.1021/je050134y
Interfacial tension between Water and n-Octane plus Sorbitan ... (274.2 to 293.2) K	https://www.doi.org/10.1021/je101094u
Liquid-liquid equilibria of ionic liquid + thiophene + ...	https://www.doi.org/10.1016/j.jct.2019.03.017
Measurements of activity coefficients at infinite dilution for organic solutes and water in N-hexylisoquinolinium thiocyanate, [HiQuin][SCN] using GLC	https://www.doi.org/10.1016/j.jct.2013.02.004
1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography	https://www.doi.org/10.1021/je800846j
Mixtures of Ethyl Alcohol with ... Activity Coefficients at Infinite Dilution of Organic Solutes and Crystal Structure	https://www.doi.org/10.1021/je800925v
Measurement of critical properties for binary and ternary mixtures containing potential gasoline additive diethyl carbonate (DEC)	https://www.doi.org/10.1016/j.fluid.2018.04.021
Liquid–liquid equilibrium of (perfluoroalkane + alkane) binary mixtures	https://www.doi.org/10.1016/j.fluid.2006.02.003
Thermodynamic Properties of Mixtures Containing Ionic Liquids: Activity Coefficients at Infinite Dilution of Organic Compounds	https://www.doi.org/10.1021/je060033f
Enthalpies of Vaporization and Vapor Pressures of Some Deuterated ... Bromide and 3-Alkylimidazolium	https://www.doi.org/10.1021/je800091s
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 8. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in 1-Propenyl-3-alkylimidazolium Bromide	https://www.doi.org/10.1021/je0503554
Activity coefficients at infinite dilution	https://www.doi.org/10.1016/j.jct.2013.10.038
Thermodynamics and activity coefficients at infinite dilution for organic solutes	https://www.doi.org/10.1016/j.jct.2017.07.012
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography	https://www.doi.org/10.1021/je800218g
Using Gas-Liquid Chromatography: extractive desulfurization using ...	https://www.doi.org/10.1016/j.fluid.2017.10.017
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate using GLC	https://www.doi.org/10.1016/j.fluid.2009.01.011
Thermophysical properties for (diethyl carbonate + p-xylene + octane) ternary system	https://www.doi.org/10.1016/j.jct.2011.07.012

Activity coefficients at infinite dilution of alkanes and alkenes in 1-alkyl-3-methylimidazolium tetrafluoroborate	https://www.doi.org/10.1016/j.fluid.2009.05.007
Viscosity Measurement of Endothermic Fuels at Temperatures from 303 K to 673 K and Pressures up to 5.00 MPa:	https://www.doi.org/10.1021/acs.jced.6b00391
Densities and Viscosities of Binary Mixtures of Ethylbenzene + N-Methyl-2-pyrrolidone, Ethylbenzene + Sulfolane, and Styrene + Octane from (303.15 to 353.15) K and Atmospheric Pressure	https://www.doi.org/10.1021/je049572f
Determination and correlation of liquid–liquid equilibria for nine binary 1-Alkyl-1-methylpyrrolidinium Tetrafluoroborate Ionic Liquids + Alkane Systems	https://www.doi.org/10.1016/j.fluid.2010.10.013
Activity Coefficients at Infinite Dilution for Organic Solutes Dissolved in Three Alkanediol Isomers at 298.15 K	https://www.doi.org/10.1021/je4001894
Abraham model correlations for describing the thermodynamic properties of dissolution in Ternary (liquid + liquid) equilibria of ionic liquids bearing short linear alkyl side chains of ether-functionalized cations and correlations:	https://www.doi.org/10.1016/j.jct.2018.05.003
Ternary (liquid + liquid) equilibria of 1-alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ionic liquids + alkane + alkanol at 298.15 K: Experiments and correlations:	https://www.doi.org/10.1016/j.fluid.2012.12.025
Joback Method	https://en.wikipedia.org/wiki/Joback_method
Liquid liquid equilibria for the ternary system water + octane + 2-butoxy-ethanol:	https://www.doi.org/10.1016/j.fluid.2015.04.019
Solubility of Mercury in Liquid Hydrocarbons and Hydrocarbon Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00173

Legend

af:	Acentric Factor
aigt:	Autoignition Temperature
ap:	Aniline Point
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hcg:	Heat of Combustion, Gross form
hcn:	Heat of Combustion, Net Form
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hsubt:	Enthalpy of sublimation at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
pc:	Critical Pressure

pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
sfust:	Entropy of fusion at a given temperature
sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tcondl:	Liquid thermal conductivity
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter